

Atomic Structures Study Guide **ANSWERS**

1. The Periodic Table is mostly made of metals.
2. The noble gases are the farthest group to the right on the Periodic Table.
3. Atoms having equal numbers of positive and negative charges make them neutral.
4. The three primary sub-atomic particles that atoms are made of are protons, neutrons & electrons.
5. Each element is made up of only one type of atom and it is represented on the Periodic Table by a symbol.
6. There are about 100 types of elements or atoms and 92 of these are occurring naturally.
7. An ion is an atom or molecule that has gained (become more negative) or lost electrons (become more positive) and thus has an electrical charge.
8. A molecule is the smallest complete part of a compound.
9. The nucleus is the center of an atom containing protons and neutrons.
10. Identify these parts of an atom:
(A) protons are found in the nucleus and have a positive charge.
(B) neutrons are found in the nucleus and have no charge; they are neutral.
(C) electrons are found outside the nucleus and have a negative charge.
11. The number of protons in an atom is also called its atomic number.
12. To find the atomic mass of an atom, you add the number of protons and neutrons in the nucleus.
13. Energy levels, shells, orbitals or clouds outside the nucleus are where you will find the electrons.
14. In the energy levels, shells, orbitals or clouds is where all chemical reactions occur and never in the nucleus.
15. The atomic number identifies the element and is the number of protons in the nucleus of the atom.
16. The atomic mass is the mass of the atom and it is the number of protons and neutrons in the nucleus.
17. The rows across the Periodic Table in which the number of energy levels in each type of atom (or element) is the same are called periods.

Physics The Atom Study Guide Answers

Jin-Ying Zhang



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