

# General Chemistry Placement Test, With Complete Solutions Updated | 2024\2024.

molecule two nonmetal atoms combined form a covalent compound called a  
diatomic molecules  $H_2, O_2, F_2$   
Matter Anything that has mass and takes up space  
Element A substance that cannot be broken down into simpler substances  
Compound A pure substance made of two or more elements CHEMICALLY combined.  
Compound Can be decomposed by chemically  
Mixture A PHYSICAL blend of two or more components  
(aq) aqueous mixture, a substance is dissolved in water  
homogeneous mixture A mixture that is so evenly mixed that you can't see all of the different parts. Ex. Air, kool-aid  
heterogeneous mixture A mixture in which different materials can be seen easily.  
Ex. soil, chocolate chip cookies  
filtration process to separate a solid "precipitate" from a liquid  
sublimation phase change from (s)  $\rightarrow$  (g)  
empirical formula A formula showing the lowest whole number ratio of atoms in a compound. Can't be reduced.  
solid A state of matter that has a definite shape and a definite volume  
liquid A form of matter that flows, has constant volume, and takes the shape of its container  
Bromine (Br) and Mercury (Hg) only two liquids on the periodic table  
gas A state of matter with no definite shape or volume  
gas will uniformly and completely fill a container  
gases physical property most all nonmetals and group 18  
physical property A characteristic of a substance that does not involve a chemical change.  
Ex. melting/boiling point, solubility, density, color, or hardness.  
properties used to identify or recognize matter (objects).  
chemical property describes how matter reacts when it changes chemically into another substances. Ex. reactivity with acid, base or with air to rust, flammability, decomposition.  
physical/phase change(s)  $\rightarrow$  (l)  $\rightarrow$  (g), diluting, crush, cut. Still same substance.  
Density A physical property. Specific to every element.  
Chemical symbol One capital letter, or a capital letter and a lower case letter together. Mass on upper left side, and atomic # on lower left side.  
Conservation of Mass Equal # of atoms on both sides of a chemical equation  
Democritus Greek philosopher that said all matter is made of tiny particles called "atomos" or atoms  
JJ Thomson used the cathode ray tube to discover electrons

# Usc Chemistry Placement Test Review

**College Research Group of Concord,  
Massachusetts**



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