## 7-1 Review and Reinforcement

## Ionic Bonding

Complete each of the following sentences by filling in the appropriate word or phrase from the list below.

negative positive crisscross noble gas electrons anion

octet ions cation electron configuration polyatomic empirical formula

- An ionic bond is an attraction between oppositely charged —ions
- Anions have a -negative ----- charge.
- An atom becomes an ion by losing or gaining \_\_\_\_electrons
- The \_octet \_\_rule states that atoms tend to gain, lose, or share electrons in order to acquire a full set of valence electrons.
- When sodium and chlorine form an ionic bond, both ions acquire the electron configuration of a(n) \_noble gas
- 6. The atoms that make up **polyatomic** \_ions are bonded together by covalent bonds.
- The \_crisscross \_\_method can be used to write the formula for an ionic compound.
- 8. The \_empirical formula of a compound denotes the ratio of ions in the compound.

If the statement is true, write "true." If it is false, change the underlined word or words to make it true. Write your answer on the line.

- losing . 9. Calcium becomes a monatomic cation by gaining two electrons.
- \_ true \_\_\_\_\_ 10. A cation has a <u>positive</u> charge.
- true . 11. A binary ionic compound contains only gag kind of cation and one kind of anion.
- \_polyatomic \_\_ 12. Monatomic ions consist of more than one atom.
- \_\_seven \_\_\_\_\_13. The Lewis dot diagram for chlorine, a group 7A element, has aix electrons.
- \_\_true (released) \_ 14. A great deal of energy is produced when an ionic compound is formed.

## **Section 1 Reinforcement Kinetic Theory Answer Key**

**M Woodhall** 

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 $, 1970\text{-}07 \qquad \textbf{Engineered Materials Abstracts} \ , 1988\text{-}07$ 

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