

## 14.3

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on  $K_a$  and  $K_b$  and vice versa.
3. Calculate the pH of weak acids and bases.

## Vocabulary

- |                        |                             |                   |
|------------------------|-----------------------------|-------------------|
| 1. autoionization      | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. ionization constant | 5. $K_a$                    | 6. $K_b$          |
|                        | 7. weak solution            |                   |

## Key Equation

- |                                     |                                     |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

## Part A Completion

Write the equation and/or the numerical value(s) for the reaction(s) indicated. Show the calculation for pH or percent ionization when required. Assume acids are monoprotic unless noted.

1. What substance is \_\_\_\_\_, a weak hydrogen ion donor?
2. \_\_\_\_\_ is the concentration of those species present in solution that equal to \_\_\_\_\_.
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20. \_\_\_\_\_ is the concentration of those species present in solution that equal to \_\_\_\_\_.

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## Section 19 2 Hydrogen Ions And Acidity Answer Key

**Disha Experts**

A red circular graphic with a gradient, appearing as a partial circle or a stylized 'C' shape, located to the right of the 'Disha Experts' text.

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