

Name _____ Class _____ Date _____

Section 2-2 Properties of Water (pages 40–43)**Key Concepts**

- Why are water molecules polar?
- What are acidic solutions? What are basic solutions?

The Water Molecule (pages 40–41)

1. Is the following sentence true or false? A water molecule is neutral. true
2. Why is a water molecule polar? There is an uneven distribution of electrons between the oxygen and hydrogen atoms.
3. Circle the letter of each sentence that is true about hydrogen bonds.
 - a. A hydrogen bond is stronger than an ionic bond.
 - b.** The attraction between the hydrogen atom on one water molecule and the oxygen atom on another water molecule is an example.
 - c. A hydrogen bond is stronger than a covalent bond.
 - d.** They are the strongest bonds that form between molecules.
4. Complete the table about forms of attraction.

FORMS OF ATTRACTION

Form of Attraction	Definition
Cohesion	<u>Attraction between molecules of the same substance</u>
Adhesion	<u>Attraction between molecules of different substances</u>

Solutions and Suspensions (pages 41–42)

5. What is a mixture? A mixture is a material composed of two or more elements or compounds that are physically mixed together but not chemically combined.
6. A mixture of two or more substances in which the molecules of the substances are evenly mixed is called a(an) solution.
7. The greatest solvent in the world is water.
8. What is a suspension? A suspension is a mixture of water and nondissolved substances that are so small they do not settle out.

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Jay K. Hackett



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