

## Section 18.2

### Reversible Reactions and Equilibrium

#### OBJECTIVES

- Describe how the *amounts of reactants and products change* in a chemical system at equilibrium.
- Identify three *stresses* that can change the equilibrium position of a chemical system.
- Explain what the value of  $K_{eq}$  indicates about the position of equilibrium.

#### Reversible Reactions

- Some reactions do not go to completion as we have assumed
    - They may be \_\_\_\_\_ – a reaction in which the conversion of reactants to products and the conversion of products to reactants occur simultaneously
  - \_\_\_\_\_ :  $2\text{SO}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{SO}_3(\text{g})$
  - \_\_\_\_\_ :  $2\text{SO}_3(\text{g}) + \text{O}_2(\text{g}) \leftarrow 2\text{SO}_2(\text{g})$
  - The two equations can be combined into one, by using a \_\_\_\_\_, which tells us that it is a reversible reaction:  
 $2\text{SO}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{SO}_3(\text{g})$
  - ✓ A chemical *equilibrium* occurs, and no net change occurs in the actual amounts of the components of the system.
  - Even though the *rates* of the forward and reverse are equal, the \_\_\_\_\_ of components on both sides may not be equal
    - An equilibrium position may be shown:
- |    |     |    |     |    |
|----|-----|----|-----|----|
| A  | B   | or | A   | B  |
| 1% | 99% |    | 99% | 1% |
- ✓ Note the emphasis of the arrows direction
  - ✓ It depends on which side is favored; almost all reactions are reversible to some extent

#### Le Chatelier's Principle

- The French chemist Henri Le Chatelier (1850-1936) studied how the equilibrium position *shifts* as a result of changing conditions
  - Le Chatelier's principle: \_\_\_\_\_
-

# Section 18 2 Reversible Reactions And Equilibrium

## Answer Key

**SB Merriam**



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