

11.1

DESCRIBING CHEMICAL REACTIONS

Section Review

Definitions

- 1. **Reactants** are substances that react.
- 2. **Products** are substances formed by a reaction.
- 3. **Chemical equation** is a symbolic representation of a reaction.

Classification

- 1. **Combination reactions**
- 2. **Decomposition reactions**
- 3. **Single-replacement reactions**



Plant A: Combustion

After this combustion reaction, how many atoms of oxygen are present in the products? Identify them in the space below.

Reaction: $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$

This combustion reaction is also known as **burning**. This term describes the oxidation of a substance, such as the **fuel**, **methane**, shown here. The fuel reacts with **oxygen** from the air to form **carbon dioxide** and **water vapor**. **Water vapor** is a gas that is **less dense** than **air**. As a result, it rises. **Carbon dioxide** is a gas that is **more dense** than **air**. As a result, it stays near the bottom of the container. **Water vapor** is composed of **hydrogen** and **oxygen** atoms. **Carbon dioxide** is composed of **carbon** and **two oxygen** atoms.

Hydrogen atoms are represented by the symbol H and oxygen atoms are represented by the symbol O . Carbon atoms are represented by the symbol C . **Hydrogen** and **oxygen** atoms are found in almost all substances. A compound containing hydrogen, oxygen, and carbon atoms is called a **carbon-containing compound**.

- 1. **Hydrogen**
- 2. **Oxygen**
- 3. **Potassium**
- 4. **Sulfur**
- 5. **Hydrochloric acid**
- 6. **Ammonium nitrate**
- 7. **Hydrogen sulfide**
- 8. **Hydrogen**
- 9. **Hydrogen**
- 10. **Hydrogen**
- 11. **Hydrogen**
- 12. **Hydrogen**
- 13. **Hydrogen**
- 14. **Hydrogen**
- 15. **Hydrogen**
- 16. **Hydrogen**
- 17. **Hydrogen**
- 18. **Hydrogen**
- 19. **Hydrogen**
- 20. **Hydrogen**

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Table of Contents Section 1 2 Describing Chemical Reactions Review And Reinforce

1. Understanding the eBook Section 1 2 Describing Chemical Reactions Review And Reinforce
 - The Rise of Digital Reading Section 1 2 Describing Chemical Reactions Review And Reinforce
 - Advantages of eBooks Over Traditional Books
2. Identifying Section 1 2 Describing Chemical Reactions Review And Reinforce
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Section 1 2 Describing Chemical Reactions Review And Reinforce
 - User-Friendly Interface
4. Exploring eBook Recommendations from Section 1 2 Describing Chemical Reactions Review And Reinforce
 - Personalized Recommendations

Section 1 2 Describing Chemical Reactions Review And Reinforce

- Section 1 2 Describing Chemical Reactions Review And Reinforce User Reviews and Ratings
 - Section 1 2 Describing Chemical Reactions Review And Reinforce and Bestseller Lists
5. Accessing Section 1 2 Describing Chemical Reactions Review And Reinforce Free and Paid eBooks
- Section 1 2 Describing Chemical Reactions Review And Reinforce Public Domain eBooks
 - Section 1 2 Describing Chemical Reactions Review And Reinforce eBook Subscription Services
 - Section 1 2 Describing Chemical Reactions Review And Reinforce Budget-Friendly Options
6. Navigating Section 1 2 Describing Chemical Reactions Review And Reinforce eBook Formats
- ePUB, PDF, MOBI, and More
 - Section 1 2 Describing Chemical Reactions Review And Reinforce Compatibility with Devices
 - Section 1 2 Describing Chemical Reactions Review And Reinforce Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Section 1 2 Describing Chemical Reactions Review And Reinforce
 - Highlighting and Note-Taking Section 1 2 Describing Chemical Reactions Review And Reinforce
 - Interactive Elements Section 1 2 Describing Chemical Reactions Review And Reinforce
8. Staying Engaged with Section 1 2 Describing Chemical Reactions Review And Reinforce
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Section 1 2 Describing Chemical Reactions Review And Reinforce
9. Balancing eBooks and Physical Books Section 1 2 Describing Chemical Reactions Review And Reinforce
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Section 1 2 Describing Chemical Reactions Review And Reinforce
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Section 1 2 Describing Chemical Reactions Review And Reinforce
- Setting Reading Goals Section 1 2 Describing Chemical Reactions Review And Reinforce
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Section 1 2 Describing Chemical Reactions Review And Reinforce
- Fact-Checking eBook Content of Section 1 2 Describing Chemical Reactions Review And Reinforce

- Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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