

1. Hydrogen gas and Iron chloride are produced.

Explanation: The Chlorine from Hydrogen chloride is displaced by the Iron fillings to undergo the following reaction.

2HCl + Fe → FeCl₂+ H₂

4. What is a balanced chemical equation? Why should a chemical equation be balanced?

Sobation:

A balanced equation is the one in which number of different atoms on both the reactant and product sides are equal. Balancing chemical equations is necessary for the reaction should obey The Law of Conservation of mass. Balancing the chemical equation has no defined method and is purely a trial-and-error attempt.

- 5. Translate the following statements into chemical equations and balance them.
- (a) Hydrogen gas combines with nitrogen to form ammonia.
- (b) Hydrogen sulphide gas burns in the air to give water and sulphur dioxide.

(c) Barium chloride reacts with aluminium sulphate to give Aluminium chloride and a precipitate of barium sulphate.

(d) Potassium metal reacts with water to give potassium hydroxide and Hydrogen gas.

Sobuttions:

- (a) Unbalanced: H₂ + N₂ → NH₃
- Balanced: 3H.+ N. -- 2NH.
- (b) Unbalanced: H₂S + O₂ → H₂O + SO₃
- Balanced: $2H.S + 3O. \rightarrow 2H.O + 2SO.$
- (c) Unbalanced:
- $BaCl_t + Al_t(SO_t) \rightarrow AlCl_t + BaSO_t$
- Balanced: $3BaCl_1 + Al_2(SO_4) \rightarrow 2AlCl_4 + 3BaSO_4$
- (d) Unbalanced: K + H₂O → KOH + H₃
- Balanced: 2K + 2H₂O → 2KOH + H₂

Balance the following chemical equations.

- (a) $HNO_2 + Ca(OH)_1 \rightarrow Ca(NO_2)_2 + H_2O$
- (b) $NnOH + H_iSO_i \rightarrow Nn_iSO_i + H_iO$
- (c) NaCl + AgNO₂ → AgCl + NaNO₂
- (d) BaCl₂ + H₂SO₄ → BaSO₄ + HCl

Solution:

(a) 2HNO_i+ Ca(OH)_i → Ca(NO_i)_i+ 2H_iO

Section 1 Chemical Changes Answers

Schyrlet Cameron, Carolyn Craig

Section 1 Chemical Changes Answers:

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