

## 7-1 Review and Reinforcement

### Ionic Bonding

Complete each of the following sentences by filling in the appropriate word or phrase from the list below.

negative  
positive  
crisscross

noble gas  
electrons  
anion

octet  
ions  
cation

electron configuration  
polyatomic  
empirical formula

1. An ionic bond is an attraction between oppositely charged ions.
2. Anions have a negative charge.
3. An atom becomes an ion by losing or gaining electrons.
4. The octet rule states that atoms tend to gain, lose, or share electrons in order to acquire a full set of valence electrons.
5. When sodium and chlorine form an ionic bond, both ions acquire the electron configuration of a(n) noble gas.
6. The atoms that make up polyatomic ions are bonded together by covalent bonds.
7. The crisscross method can be used to write the formula for an ionic compound.
8. The empirical formula of a compound denotes the ratio of ions in the compound.

If the statement is true, write "true." If it is false, change the underlined word or words to make it true. Write your answer on the line.

- |                        |                                                                                            |
|------------------------|--------------------------------------------------------------------------------------------|
| <u>losing</u>          | 9. Calcium becomes a monatomic cation by <u>gaining</u> two electrons.                     |
| <u>true</u>            | 10. A cation has a <u>positive</u> charge.                                                 |
| <u>true</u>            | 11. A binary ionic compound contains only <u>one</u> kind of cation and one kind of anion. |
| <u>polyatomic</u>      | 12. <u>Monatomic</u> ions consist of more than one atom.                                   |
| <u>seven</u>           | 13. The Lewis dot diagram for chlorine, a group 7A element, has <u>six</u> electrons.      |
| <u>true (released)</u> | 14. A <u>great deal of energy is produced</u> when an ionic compound is formed.            |

# Section 1 Reinforcement 7 Answers

**H Kauffman**



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## **Table of Contents Section 1 Reinforcement 7 Answers**

1. Understanding the eBook Section 1 Reinforcement 7 Answers
  - The Rise of Digital Reading Section 1 Reinforcement 7 Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Section 1 Reinforcement 7 Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Section 1 Reinforcement 7 Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Section 1 Reinforcement 7 Answers
  - Personalized Recommendations
  - Section 1 Reinforcement 7 Answers User Reviews and Ratings
  - Section 1 Reinforcement 7 Answers and Bestseller Lists
5. Accessing Section 1 Reinforcement 7 Answers Free and Paid eBooks
  - Section 1 Reinforcement 7 Answers Public Domain eBooks

- Section 1 Reinforcement 7 Answers eBook Subscription Services
- Section 1 Reinforcement 7 Answers Budget-Friendly Options
- 6. Navigating Section 1 Reinforcement 7 Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Section 1 Reinforcement 7 Answers Compatibility with Devices
  - Section 1 Reinforcement 7 Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Section 1 Reinforcement 7 Answers
  - Highlighting and Note-Taking Section 1 Reinforcement 7 Answers
  - Interactive Elements Section 1 Reinforcement 7 Answers
- 8. Staying Engaged with Section 1 Reinforcement 7 Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Section 1 Reinforcement 7 Answers
- 9. Balancing eBooks and Physical Books Section 1 Reinforcement 7 Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Section 1 Reinforcement 7 Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Section 1 Reinforcement 7 Answers
  - Setting Reading Goals Section 1 Reinforcement 7 Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Section 1 Reinforcement 7 Answers
  - Fact-Checking eBook Content of Section 1 Reinforcement 7 Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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