

Physical Science – 2nd Semester – Final Exam Study Guide – (2012–2013)

1. Physical Science Basics

- What tool(s) would you want to use to find the:
 - Mass of an object? triple beam balance Basic SI Unit? kg
 - Volume of an object? ruler or water displacement method Basic SI Unit? cm³ or ml
 - Length of an object? Ruler Basic SI Unit? meter
 - Density of an object? graduated cylinder Basic SI Unit? ml
- What is the formula for calculating density? mass/volume
 - What is the density of a cube that is 25 g and has a volume of 3 cm³? 8g/cm³
 - Compare the density of solid H₂O to liquid H₂O, and explain what observable effect this has.
solid H₂O less dense than liquid H₂O

2. Chemistry Basics

- Element (def.) substance made up of one kind of atom
- Atom (def.) basic building block of matter
- Compound (def.) a molecule that contains atoms or molecules of two or more different elements
- Molecule (def.) a combination of two or more of the same or different atoms that are bonded together

3. Model of the Atomic Evolution of the Atomic Theory

Atoms are so small that many scientists throughout history have created models to describe them.

Why have atomic theories changed over time? Better technology came around.

Scientist	Year	Major Discovery/ Contribution	Name of atomic model
John Dalton	1808	Each element is made up of small, indivisible particles called atoms. Different elements have atoms of different masses.	Billiard Ball Model
J.J. Thomson	1897	Negatively charged electron	Plum Pudding Model
Niels Bohr	1913	Electrons move <u>fixed orbits</u> around the atom's nucleus in specific layers, or <u>electron shells</u> . Atoms absorb or give off energy when the electrons move from one shell (<u>energy level</u>) to another.	Bohr Model
Ernest Rutherford	1909	1909 /Based on idea that positive reflects positive/ atom mostly empty space; electrons orbit randomly around positively charged nucleus	Planetary Model

4. Periodic Table

- The current periodic table is arranged by increasing atomic number.
- Reading the Periodic Table
 - An element's properties can be predicted from its location in the periodic table.
 - The atomic number represents the number of protons.
 - The atomic mass represents the number of protons and neutrons.

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Modoc Press



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Title Index, ERIC ED Accessions File ERIC Processing and Reference Facility,1966 **United States Armed Forces Institute Catalog** United States Armed Forces Institute,1961 **The Macmillan Guide to Correspondence Study** Modoc Press,1996 **An Introduction to Physical Science** James T. Shipman,Jerry L. Adams,Jerry D. Wilson,1979

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Table of Contents Second Semester Physical Science Final Review Guide

1. Understanding the eBook Second Semester Physical Science Final Review Guide
 - The Rise of Digital Reading Second Semester Physical Science Final Review Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Second Semester Physical Science Final Review Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Second Semester Physical Science Final Review Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Second Semester Physical Science Final Review Guide
 - Personalized Recommendations
 - Second Semester Physical Science Final Review Guide User Reviews and Ratings
 - Second Semester Physical Science Final Review Guide and Bestseller Lists

5. Accessing Second Semester Physical Science Final Review Guide Free and Paid eBooks
 - Second Semester Physical Science Final Review Guide Public Domain eBooks
 - Second Semester Physical Science Final Review Guide eBook Subscription Services
 - Second Semester Physical Science Final Review Guide Budget-Friendly Options
6. Navigating Second Semester Physical Science Final Review Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Second Semester Physical Science Final Review Guide Compatibility with Devices
 - Second Semester Physical Science Final Review Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Second Semester Physical Science Final Review Guide
 - Highlighting and Note-Taking Second Semester Physical Science Final Review Guide
 - Interactive Elements Second Semester Physical Science Final Review Guide
8. Staying Engaged with Second Semester Physical Science Final Review Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Second Semester Physical Science Final Review Guide
9. Balancing eBooks and Physical Books Second Semester Physical Science Final Review Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Second Semester Physical Science Final Review Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Second Semester Physical Science Final Review Guide
 - Setting Reading Goals Second Semester Physical Science Final Review Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Second Semester Physical Science Final Review Guide
 - Fact-Checking eBook Content of Second Semester Physical Science Final Review Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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