Name:

Neutralization Reactions Worksheet Reading: Cycle 6 Chemistry II Lesson 4

Neutralization Reaction: Acid + Base \rightarrow Salt + Water H * from acid and OH * from base combine to form H $_2$ O. Leftovers form 'salt'. Remember in salt, write metal first! KCl not CIK!

Example 1: $H + N_0OH - H_2O + N_0I$

H and OH combine to form water. One atom of I & one of Na leftover. Na is a metal so write NaI (not INa)

Example 2: __HF + _ Ca(OH)₂ -- ___ + ___

For complete neutralization, must have equal number of H and OH. Therefore we need 2 HF for 1 Ca(OH)₂

_2_HF + _1_Ca(OH)₂ -- ____ +____

Two H and two OH make two waters. Think of prom! Two boys and two girls make two couples.

_2_HF + _1_Ca(OH)₂ → 2 H₂O +

What's left? 2 F atoms and 1 Ca atom form the 'salt'. Remember that Ca is the metal, so it goes first.

 $2 \text{ HF} + 1 \text{ Ca(OH)}_2 - 2 \text{ H}_2 \text{O} + \text{CaF}_2$

Safety check: is CaF₂ a good formula? P.T. tells you ions are Card and F, so total charge is 0 like it should be.

Complete & balance these neutralization reactions. Remember that H₂O = HOH!

 $H_2SO_4 + Sr(OH)_2 \rightarrow __H_2O + SrSO_4$ $_Ca(OH)_2 + __H_3PO_4 \rightarrow __H_2O + __Ca_5(PO_4)_2$ $_HBr + Ba(OH)_2 \rightarrow __H_2O +$

 $Zn(OH)_2 + _HNO_3 - __ + Zn(NO_3)_2$

Al(OH)₃ + __HCl -- ____ + ____

Complete and balance the following equations representing neutralization reactions:

a) _CsOH + H₂SO₃ - _H₂O + Cs₂CO₃

b) _HF + Mg(OH)₂ -- _H₂O + _____

c) $_HNO_3$ + $Al(OH)_3$ - $_H_2O$ + $Al(NO_3)_3$

d) HCl + KOH - ____ + _____

e) HBrO₃ + LiOH - H₂O + _____

Section 19 4 Neutralization Reactions Worksheet Answers

Wolfgang Guggemos

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Canadian Journal of Medical Technology ,1979

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