

Chapter 2: Water

2.1 Water Weak Interactions in Aqueous Systems

Water's Qualities:

-Unusual properties of water caused by Hydrogen Bonds:

- 1) High Surface tension – beads up
- 2) High heat capacity
- 3) High Heat of Vaporization – a lot of energy is needed for water to boil
- 4) Expands as it freezes – the solid forms of water (ice) floats instead of sinks like most

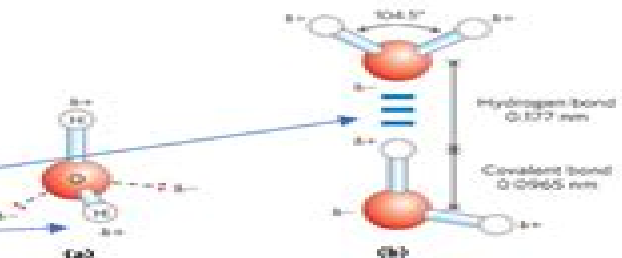
Inter/Intramolecular Forces 1-HYDROGEN BONDING

-Hydrogen bonding occurs when the hydrogen atoms of one molecule is attracted to another molecule that has Nitrogen, Oxygen, and Sulfur (NOS).

- Oxygen is the most electronegative followed by sulfur and Nitrogen.

-Picture to the right: Water Molecules

- O is more electronegative than H
- This creates a POLAR covalent bond between H and O
- The H-Bond occurs between water molecules
- O has a partial negative (-) charge
- H has a partial positive (+) charge



-Polar function groups form H-Bonds with one another and water



-Angle and distance affect the strength of H-Bonds



-we will see H-Bonds occur between:

- Hydroxyl group of alcohol and water
- Carbonyl group of a ketone and water
- between peptide groups and peptides
- complementary bases of DNA

Between the hydroxyl group of an alcohol and water



Between the carbonyl group of a ketone and water



Between peptide groups in polypeptides



Between complementary bases of DNA



Section 2 Chem Study Guide Answers

Bruce Wingerd, Patty Bostwick Taylor



Section 2 Chem Study Guide Answers:

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