

# Relative Strengths of **Acids** and **Bases** in Aqueous Solutions at 25°C

Name of Acid	Acid		Base	$K_a$
Hydriodic	HI	→	$H^+ + I^-$	$>> 1$
Hydrobromic	HBr	→	$H^+ + Br^-$	$>> 1$
Hydrochloric	HCl	→	$H^+ + Cl^-$	$>> 1$
Sulfuric	$H_2SO_4$	→	$H^+ + HSO_4^-$	$> 1$
Nitric	$HNO_3$	→	$H^+ + NO_3^-$	$> 1$
Hydronium Ion	$H_3O^+$	⇌	$H^+ + H_2O$	$5.50 \times 10^{-1}$
Oxalic	$H_2C_2O_4$	⇌	$H^+ + HC_2O_4^-$	$5.90 \times 10^{-2}$
Sulfurous	$H_2SO_3$	⇌	$H^+ + HSO_3^-$	$1.41 \times 10^{-2}$
Hydrogen Sulfate Ion	$HSO_4^-$	⇌	$H^+ + SO_4^{2-}$	$1.20 \times 10^{-2}$
Phosphoric	$H_3PO_4$	⇌	$H^+ + H_2PO_4^-$	$7.52 \times 10^{-3}$
Nitrous	$HNO_2$	⇌	$H^+ + NO_2^-$	$5.62 \times 10^{-4}$
Hydrofluoric	HF	⇌	$H^+ + F^-$	$3.53 \times 10^{-4}$
Formic	$HCO_2H$	⇌	$H^+ + HCO_2^-$	$1.78 \times 10^{-4}$
Benzoic	$C_6H_5CO_2H$	⇌	$H^+ + C_6H_5CO_2^-$	$6.46 \times 10^{-5}$
Acetic	$CH_3CO_2H$	⇌	$H^+ + CH_3CO_2^-$	$1.76 \times 10^{-5}$
Carbonic	$H_2CO_3$	⇌	$H^+ + HCO_3^-$	$4.30 \times 10^{-7}$
Hydrogen Sulfite Ion	$HSO_3^-$	⇌	$H^+ + SO_3^{2-}$	$1.02 \times 10^{-7}$
Dihydrogen Phosphate Ion	$H_2PO_4^-$	⇌	$H^+ + HPO_4^{2-}$	$6.23 \times 10^{-8}$
Boric	$H_3BO_3$	⇌	$H^+ + H_2BO_3^-$	$5.79 \times 10^{-10}$
Ammonium Ion	$NH_4^+$	⇌	$H^+ + NH_3$	$5.64 \times 10^{-10}$
Hydrogen Carbonate Ion	$HCO_3^-$	⇌	$H^+ + CO_3^{2-}$	$5.61 \times 10^{-11}$
Monohydrogen Phosphate Ion	$HPO_4^{2-}$	⇌	$H^+ + PO_4^{3-}$	$2.20 \times 10^{-13}$
Water	$H_2O$	⇌	$H^+ + OH^-$	$1.00 \times 10^{-14}$

Strong Acids

Weak Acids

Weakest Acids

Weakest Bases

Weak Bases

Stronger Bases

All reactions assume water as a reactant.



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$$K_a = [H^+][A^-] / [HA]$$

$$pK_a = -\log K_a$$

$$pH = -\log [H^+]$$

## Section 2 Study Guid Strength Of Acid And Bases

**Leo J. Malone**



## **Section 2 Study Guid Strength Of Acid And Bases:**

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