

## Rates of reactions

Specification reference	Learning objectives	
6.3.1.1a	Given your experiment with reactive metals state all important factors, except concentration, which influence the rate of reaction.	<input type="checkbox"/>
6.3.1.1b	Given your data determine a trend from your experimental data?	<input type="checkbox"/>
6.3.1.1c	Given your data determine the equation from a graph of the form $y = kx^n$ (n = 1, 2 or 3) using logarithms of 10, 15 or 20?	<input type="checkbox"/>
6.3.1.1d	Given your knowledge that rate constants, k, are constant quantities from a rate equation, including determination of units?	<input type="checkbox"/>
6.3.1.1e	Given your determination of the order of a rate equation for a reaction from the change in a graph?	<input type="checkbox"/>
6.3.1.1f	Given your knowledge explain the factors from the rate equation of a reaction?	<input type="checkbox"/>
6.3.1.1g	Given your knowledge interpret half-life, $t_{1/2}$ , from a concentration-time graph?	<input type="checkbox"/>
6.3.1.1h	Given your determination the rate constants, k, for a first order reaction from the intercept (half-life, $t_{1/2}$ ) using the relationship: $k = \ln 2/t_{1/2}$ ?	<input type="checkbox"/>
6.3.1.1i	Given your determination of the order of a reaction from a graph?	<input type="checkbox"/>
6.3.1.1j	Given your determination the rate constants for a first order reaction from the gradient?	<input type="checkbox"/>
6.3.1.1k	Given your knowledge that two half-lives equals an effective pseudo first-order reaction, how far does this allow you to extend the method used by first-order determination?	<input type="checkbox"/>
6.3.1.1l	Given your knowledge rate constants include temperature with the rate determining step but is half-life unaffected?	<input type="checkbox"/>

## Section 18 Rates Of Reaction Answer Key

**David W. Oxtoby, H. Pat Gillis, Laurie J.  
Butler**



## Section 18 Rates Of Reaction Answer Key:

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