

Unit 9 Worksheet: St. Michael's Concentration

Part 9: Describing Molar Concentrations

1. Sodium chloride was dissolved in water to produce a 1.5M solution. Suppose you had three separate beakers, each about the size of a beaker, that weighed as follows: **1.0000 g, 1.0000 g, and 1.0000 g.**
What percentage of each beaker (1.0, 1.0, and 1.0) is made up of sodium chloride, if possible, and the percentage of water, if possible, in each beaker?
2. A 10.0 g mixture of potassium chloride is dissolved in enough water to make 100.0 mL of solution. Calculate the mass percent of potassium chloride.

$$\text{NaCl} = \frac{1.0000 \text{ g}}{1.0000 \text{ g}} \times 100\% = 100\%$$

$$\text{H}_2\text{O} = \frac{1.0000 \text{ g}}{1.0000 \text{ g}} \times 100\% = 100\%$$

3. Find the molarity of a solution made from 25.0 g of CaCl_2 dissolved in enough water to make 0.5 L.

$$\text{CaCl}_2 = \frac{25.0 \text{ g}}{0.5 \text{ L}} \times \frac{1 \text{ mol}}{110.98 \text{ g}} = 0.45 \text{ M}$$

$$\text{H}_2\text{O} = \frac{1.0000 \text{ g}}{1.0000 \text{ g}} \times 100\% = 100\%$$

4. An unknown liquid solution is "diluted" and reduced to produce a 0.25 g of solution. Suppose the percentage of water is calculated as 10.0%.

$$\text{H}_2\text{O} = \frac{1.0000 \text{ g}}{1.0000 \text{ g}} \times 100\% = 100\%$$

$$\text{H}_2\text{O} = \frac{1.0000 \text{ g}}{1.0000 \text{ g}} \times 100\% = 100\%$$

5. Suppose three separate beakers are made to contain 10.0 g of CaCl_2 in 100.0 mL of solution. Calculate the molarity of each beaker.

$$\text{CaCl}_2 = \frac{10.0 \text{ g}}{100.0 \text{ mL}} \times \frac{1 \text{ mol}}{110.98 \text{ g}} = 0.09 \text{ M}$$

Unit 11 Worksheet 3 Molar Concentration

David Baker

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author's name.

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