

Equilibrium - Unit Test Review

Chapter 13 (Chemical Equilibrium)

1. Equilibrium Constant

Equilibrium Constant (K): A constant (similar to a ratio) that represents the concentrations/partial pressures of products vs. reactants at equilibrium. While there is only one equilibrium constant at a given temperature, there are an infinite number of concentrations (equilibrium positions) that satisfy it. The equations are:

$$K = \frac{[C]^c[D]^d}{[A]^a[B]^b}$$

$$K_p = K_c(RT)^{\Delta n}$$

Note that:

- If the reaction is reversed, the new K-value is $1/K$
- If the balanced equation is multiplied by a factor, K is raised by that power
- If Q (the reaction quotient) is larger than K, equilibrium favours reactants - if Q is smaller, equilibrium favours products

2. Solving Equilibrium Problems (ICE Table)

- Write out the balanced chemical equation and set up an ICE table (Initial, Change and Equilibrium concentration/pressures) - cross out solids or liquids!
- Write out initial concentrations, and then put the change in concentration as an unknown (x multiplied by the coefficient in the balanced chemical equation - +x if the substance is being formed, -x if it is being used up)
- Set equilibrium concentration to (initial - change): solve for x or substitute equilibrium concentration expressions into the equation for k to solve

<p>Note that liquids/solids DO NOT have concentrations or pressures!</p> <p>Represented as "x" and multiplied by a factor from the coefficient</p> <p>Becomes a mathematical expression to sub into the equilibrium expression</p>		$K_{eq} = P_{H_2O} \times 2 P_{HF_{aq}}$			
Initial concentration (mol/L)		[]	[]	0	[]
Change in concentration (mol/L)		-x	-x	+2x	
Equilibrium concentration (mol/L)		[]-x	[]-x	2x	

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