#### Equilibrium - Unit Test Review

#### Chapter 13 (Chemical Equilibrium)

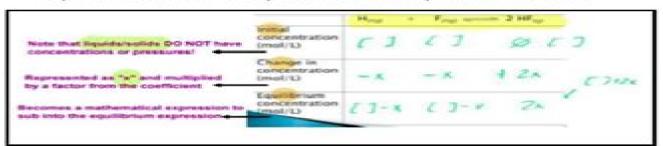
Equilibrium Constant

Equilibrium Constant (K): A constant (similar to a ratio) that represents the concentrations/partial pressures of products vs. reactants at equilibrium. While there is only one equilibrium constant at a given temperature, there are an infinite number of concentrations (equilibrium positions) that satisfy it. The equations are:

K ~ [C]~[D]~/[A]~[B]\*\* K<sub>P</sub> ~ K<sub>C</sub>(RT)\*\*

Note that:

- a. If the reaction is reversed, the new K-value is 1/K
- b. If the balanced equation is multiplied by a factor, K is raised by that power
- If Q (the reaction quotient) is larger than K, equilibrium favours reactants if
   Q is smaller, equilibrium favours products
- Solving Equilibrium Problems (ICE Table)
- Write out the balanced chemical equation and set up an ICE table (Initial, Change and Equilibrium concentration/pressures) - cross out solids or liquids!
- Write out initial concentrations, and then put the change in concentration as an unknown (x multiplied by the coefficient in the balanced chemical equation - +x if the substance is being formed, -x if it is being used up)
- Set equilibrium concentration to (initial change): solve for x or substitute equilibrium concentration expressions into the equation for k to solve



# **Unit 1test Review Equilibrium**

J Ma

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