

## CHEMISTRY 30 COURSE REVIEW

### Unit 1 – Thermochemistry

Types of Energy changes in a chemical system – temperature, phase, chemical, and nuclear

- are they potential or kinetic
- their formulas
- what type of bonding is changing

#### Terms

Thermodynamics

Law of conservation of energy

First Law of Thermodynamics

Potential energy diagram – activation energy, catalyst effect

System – open, closed, isolated

Joule/Kilojoule

Specific heat capacity ( $c = \text{J/g} \cdot ^\circ\text{C}$ )

Calorimeter heat capacity ( $C = \text{kJ}^\circ\text{C}$ )

Endothermic (+)

Exothermic (-)

Standard molar enthalpy of reaction or just Enthalpy of Reaction-new way of writing it ( $\Delta_{\text{m}}H^\circ_{\text{rxn}}$ )

Calorimetry (one side negative!) Heat loss=Heat gain

Calorimeter (simple and bomb)

Hess's Law of Heat Summation – Heat of Formation method

#### Formulas

$q$  or  $\Delta H = mc\Delta t$

where  $q$  = heat energy (J)

$m$  = mass (g)

$c$  = specific heat capacity ( $\text{J/g} \cdot ^\circ\text{C}$ )

$\Delta t$  = change in temperature ( $t_{\text{f}} - t_{\text{i}}$ )

$\Delta H = nH$

Where  $\Delta H_{\text{rxn}}$  = enthalpy of reaction (J)

$n$  = number of moles (mol)

$\Delta_{\text{m}}H$  = molar enthalpy (kJ/mol)

$\Delta H_{\text{rxn}} = \Sigma(nH)_{\text{reactants}} - \Sigma(nH)_{\text{products}}$  to find the molar heat of reaction

\*note: be able to use Hess's Law adding reactions as well as using the above formula

# Unit 1 thermochemistry Review

**Robert K. Wismer**



## Unit 1thermochemistry Review:

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