

Unit 3 Worksheet 4 – Quantitative Energy Problems Part 2

Energy constants (H₂O)

334 J/g	Heat of fusion (melting or freezing) H_f
2260 J/g	Heat of vaporization (evaporating or condensing) H_v
2.1 J/g°C	Heat capacity (c) of solid water
4.18 J/g°C	Heat capacity (c) of liquid water

For each of the problems sketch a warming or cooling curve to help you decide which equation(s) to use to solve the problem. Keep a reasonable number of sig figs in your answers.

1. How much energy must be absorbed by a 150 g sample of ice at 0.0 °C that melts and then warms to 25.0°C?
2. Suppose in the Icy Hot lab that the burner transfers 325 kJ of energy to 450 g of liquid water at 20.°C. What mass of the water would be boiled away?
3. A 12oz can of soft drink (assume $m = 340$ g) at 25°C is placed in a freezer where the temperature is - 12 °C. How much energy must be removed from the soft drink for it to reach this temperature?

Unit Worksheet Quantitative Energy Problems

Sebastian Brünink



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Table of Contents Unit Worksheet Quantitative Energy Problems

1. Understanding the eBook Unit Worksheet Quantitative Energy Problems
 - The Rise of Digital Reading Unit Worksheet Quantitative Energy Problems
 - Advantages of eBooks Over Traditional Books
2. Identifying Unit Worksheet Quantitative Energy Problems
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Unit Worksheet Quantitative Energy Problems
 - User-Friendly Interface
4. Exploring eBook Recommendations from Unit Worksheet Quantitative Energy Problems
 - Personalized Recommendations
 - Unit Worksheet Quantitative Energy Problems User Reviews and Ratings
 - Unit Worksheet Quantitative Energy Problems and Bestseller Lists
5. Accessing Unit Worksheet Quantitative Energy Problems Free and Paid eBooks

- Unit Worksheet Quantitative Energy Problems Public Domain eBooks
 - Unit Worksheet Quantitative Energy Problems eBook Subscription Services
 - Unit Worksheet Quantitative Energy Problems Budget-Friendly Options
6. Navigating Unit Worksheet Quantitative Energy Problems eBook Formats
- ePub, PDF, MOBI, and More
 - Unit Worksheet Quantitative Energy Problems Compatibility with Devices
 - Unit Worksheet Quantitative Energy Problems Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Unit Worksheet Quantitative Energy Problems
 - Highlighting and Note-Taking Unit Worksheet Quantitative Energy Problems
 - Interactive Elements Unit Worksheet Quantitative Energy Problems
8. Staying Engaged with Unit Worksheet Quantitative Energy Problems
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Unit Worksheet Quantitative Energy Problems
9. Balancing eBooks and Physical Books Unit Worksheet Quantitative Energy Problems
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Unit Worksheet Quantitative Energy Problems
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Unit Worksheet Quantitative Energy Problems
- Setting Reading Goals Unit Worksheet Quantitative Energy Problems
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Unit Worksheet Quantitative Energy Problems
- Fact-Checking eBook Content of Unit Worksheet Quantitative Energy Problems
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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