

Name: _____

Date: _____, Period: _____

Thermodynamics Review Sheet

- 1) When throwing a ball straight up into the air, at what point in its trajectory does the ball have the highest amount of gravitational potential energy? What about the highest amount of kinetic energy?
- 2) Calculate the change in internal energy of the system and determine whether the process is endothermic or exothermic.
 - a) A balloon is heated by adding 500 J of heat. It expands, doing 422 J of work on the atmosphere.
 - b) A chemical reaction releases 0.05 kJ of heat and does no work on the surroundings.
- 3) Consider the following reaction:



- a) Is heat absorbed or released in the course of this reaction? How much, in kJ?
 - b) Calculate the amount of heat transferred when 45.0 g of $\text{CH}_3\text{COH(g)}$ is decomposed by this reaction.
 - c) For a given sample of CH_3OH , the enthalpy change on reaction is 15.5 kJ. How many grams of hydrogen gas are produced from decomposing this sample of CH_3OH ?
- 4) Using the following reactions and subsequent ΔH values:



Calculate ΔH for the reaction below. (Hint = You'll need to multiply some equations to cancel things out.)



- 5) In general, how can you tell if the standard enthalpy of formation for a substance will be zero?
- 6) Calculate the standard enthalpy of formation for the following reaction:



- 7) When a 9.55-g sample of solid NaOH is added to 100.0 g of water in a calorimeter, the temperature rises from 23.4°C to 47.4°C. Assuming that the specific heat of the solution is 4.18 J/g-°C, calculate ΔH (in kJ/mol NaOH) for the reaction. (Hint: What is the mass of the solution after NaOH is added to it?)

Unit 12 Thermochemistry Test Review Answers

Richard Bailey



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