

## Unit 3 - Periodicity

### Trends in physical properties

#### Effective Nuclear Charge

- The nuclear charge of an atom is given by its atomic number. (e.g. Oxygen is +8)
- The outer electrons of the atom **do not experience the full attraction** of the nuclear charge because they are **shielded from the nucleus because of the repulsion by the inner electrons** and so the effective nuclear charge is lesser than the nuclear charge.
- Nuclear charge increases across a period, but remains constant down a group

#### Atomic Radius

- the distance from the nucleus to the outermost electrons of an atom.
- Atomic radius decreases across a period, as attraction between the nucleus and outermost electrons increases, and increases down a group, as the number of shells increases

#### Ionic Radius

- Cations: As there are less electrons than protons, the outer shell is lost and the ionic radius of a cation is lesser than the atomic radius of the atom.
  - e.g. Na has a atomic radius of  $186 \times 10^{-12}$  while  $\text{Na}^+$  has a ionic radius of  $98 \times 10^{-12}$
- Anions: As there are more electrons than protons, there is more electron repulsion and less attraction, so the ionic radius is more than the atomic radius of the atom.
  - e.g. Cl has an atomic radius of  $99 \times 10^{-12}$ , but  $\text{Cl}^-$  has an ionic radius of  $181 \times 10^{-12}$

#### Ionization Energy

- The first ionization energy of an element is the energy required to remove one mole of electrons from one mole of gaseous atoms in their ground state.
- Increases across a period as atomic bonds are stronger, and decreases down a group as there are more outshells and attraction is weaker.

#### Exceptions

- Groups 2 and 3
  - Elements of group 3 have a lower lower ionization energy than elements of Group 2.
  - Elements of Group 3 have the electron configuration of  $s^2p^1$ , which means that the electron removed is from the p orbital, which is unpaired, making it easier to remove.
  - However, elements of Group 2 have an electron configuration of  $s^2$ , which is a complete orbital and when an electron is removed, it is removed from this orbital, and it is relatively harder to remove the electron.
- Groups 15 and 16
  - Elements of Group 16 have a lower ionization energy compared to elements of Group 15.

# Unit 3 Electron Arrangement And Periodicity Study Guide Answers

**Mr. Rohit Manglik**



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