

Atomic Structure

Subatomic Particles in Neutral Atoms

The table below contains information about several isotopes. Use the information given to fill in the blanks. Assume all atoms are neutral.

Isotope Name	Neutrons by symbol	Atomic Number	Mass Number	# of Protons	# of Neutrons	# of Electrons
1. carbon-12						
2. _____		26	56			
3. _____				8	10	
4. _____	$^{235}_{92}\text{U}$					

Subatomic Particles in Ions

The table below contains information about several ions. Use the information given to fill in the blanks.

Element Name	Ions by symbol	Atomic Number	Mass Number	# of Protons	# of Neutrons	# of Electrons
5. _____	Cl^{-}				18	
6. silver			107			46
7. _____				8	8	10
8. _____	Al^{3+}		27			

Average Atomic Mass

- Calculate the average atomic mass for neon if the abundance in nature is 90.5% neon-20 (20.0 amu), 0.3% neon-21 (21.0 amu), and 9.2% neon-22 (22.0 amu).
- Calculate the average atomic mass of silver if 1.3 out of 100 atoms are silver-107 and 98.7 out of 100 atoms are silver-109.

Regents Review Atomic Structure Answers

Siddappa N.Byrareddy



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