

CHAPTER 7

THERMOCHEMISTRY SHORT QUESTION AND ANSWERS

Q.1 Define the following terms and give three examples of each.

Ans.

(i) System:

The substance which is under experiment or under observation is called as system.

Examples:

(i) $\text{Pb}(\text{NO}_3)_2$ in decomposition of $\text{Pb}(\text{NO}_3)_2$.

(ii) Zn and CuSO_4 solution, the reaction mixture in the vessel.

(iii) CaCO_3 in thermal decomposition of CaCO_3

(ii) Surroundings:

Everything around the system which is not a part of system is called surroundings.

For example

During the reaction between Zn and CuSO_4 solution vessel and air etc are surroundings.

(iii) State function:

A macroscopic property of a system which has some definite value for initial and final state and independent of the path followed e.g.

(i) Pressure (ii) Temperature (iii) Internal energy.

Note: Heat is not a state function.

Q.2 Describe the units of energy.

Ans.

Mostly Joule and calorie are used for the measurement of energy.

Calorie:

The amount of energy required to raise the temperature of one gram of water from 14.5°C to 15.5°C is called one calorie.

Joule:

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Ted Lister, Colin Osborne



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