

SI Chemistry

Chemical Reactions Lab

Purpose: to become more familiar with different types of chemical reactions by conducting several reactions, noting the results, and discussing observations with colleagues.

Procedure: At each of the nine stations, there are directions for you to follow-**including directions for waste disposal – FOLLOW ALL DIRECTIONS CAREFULLY!**

You should visit only six of the nine stations. You **MUST** visit stations 6, 8 and 9. You will then choose one station from stations 1 and 7, one station from stations 2 and 5 and one station from stations 3 and 4. You do not need to move through the stations in any particular order. You **MUST** wear your safety goggles at **ALL TIMES!**

At **each station**, answer the following questions in your lab notebook or on your tablet:

- a. How did you know a chemical reaction took place?
- b. Formula of reactant(s)
- c. Formula of product(s)
- d. Type of chemical reaction
- e. Write the balanced equation for the reaction (including all physical states).

Station 1: Get a strip of magnesium ribbon (about 2.5 cm long). Holding it with forceps, ignite the ribbon in a Bunsen burner flame. **Do not look directly at the burning magnesium!** **Looking directly at the burning magnesium could damage your eyes.** Hold the burning ribbon over a glass plate so that no burning magnesium lands on the lab bench.

Word equation: Magnesium reacts with oxygen to produce magnesium oxide.

Station 2: Place 10 drops of silver nitrate into a small test tube. Add 10 drops of sodium chloride to the same test tube.

Word equation: Aqueous silver nitrate reacts with aqueous sodium chloride to produce solid silver chloride and aqueous sodium nitrate.

Si Chemistry Chemical Reaction Lab Introduction

Sebastian Brünink



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