

SOLUBILITY CURVES

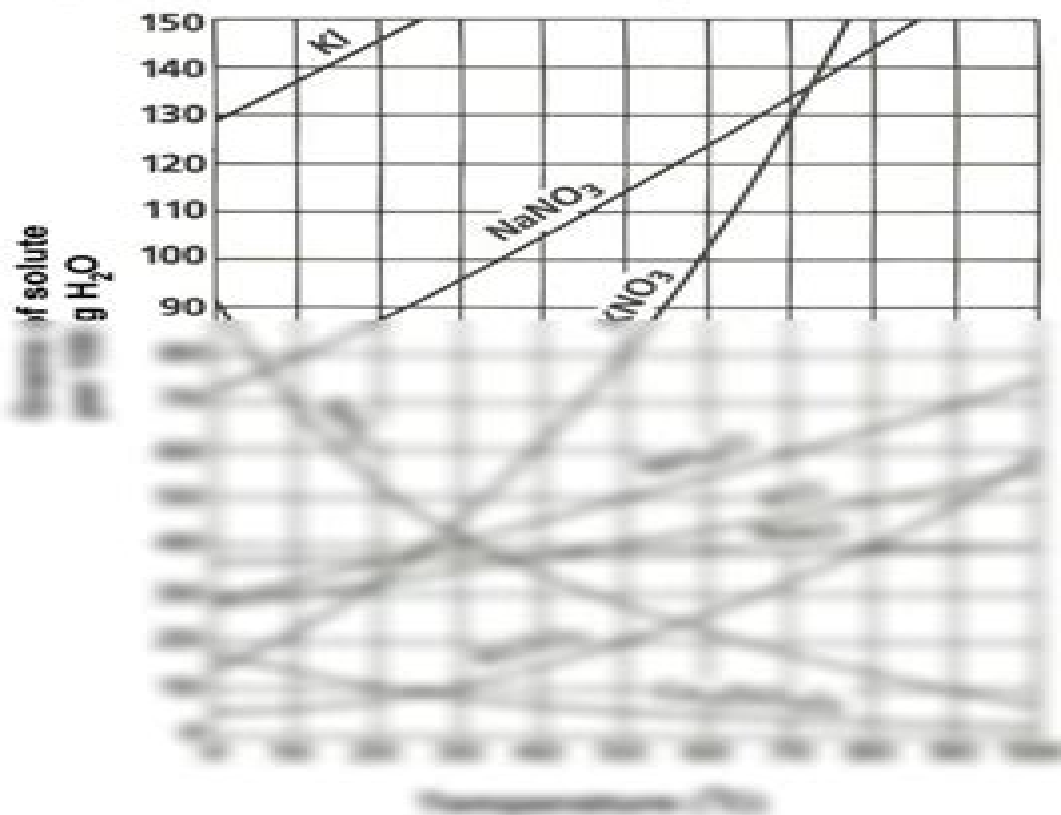
- The lines **on** a solubility curve indicate the concentration of a **saturated** solution – the maximum amount of solute that will dissolve at that specific temperature.
- Values on the graph **beneath** a curve represent **unsaturated** solutions; more solute could be dissolved at that temperature.
- A **supersaturated** solution would have a value **above** the solubility curve at a specific temperature.

Note that since density $\text{H}_2\text{O} = 1.00 \text{ g/mL}$ and since $1 \text{ mL} = 1 \text{ cm}^3$:

$$*** 100 \text{ g H}_2\text{O} = 100 \text{ mL H}_2\text{O} = 100 \text{ cm}^3 \text{ H}_2\text{O}***$$

Example: Find the curve for KClO_3 .

At 30°C approximately 10.0 g of KClO_3 will dissolve in 100g (or 100mL or 100 cm^3) of water. If the temperature is increased to 80°C , approximately 40.0 g of the substance will dissolve in 100g of water.



The most common solubility curve is for sodium chloride, which shows a very slight increase in solubility with temperature. The solubility of most salts increases with temperature, but the rate of increase varies greatly between different salts.

Key points to remember:

- Solubility is a property of a substance that describes how much of it can dissolve in a given amount of solvent at a specific temperature.
- The solubility of most salts increases with temperature, but the rate of increase varies greatly between different salts.
- Some salts, like sodium chloride, have a very slight increase in solubility with temperature.
- Other salts, like potassium nitrate, have a very sharp increase in solubility with temperature.

Solutions Solubility Curves Key

Arturo Cuomo

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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