(b) Use Eq. 1.5-12 to get: (Hen Th = 373.2°K)

At 273.2°K  $\mu = (2.21 \times 10^{-4}) \exp(38)(373.2)/(273.2) = 4.0 \times 10^{-2} q cm^2 m^2$ At 373.2°K  $\mu = (2.12 \times 10^{-4}) \exp (3.8) = 9.5 \times 10^{-5} \text{ g cm}^{-1} \text{ sec}^{-1}$ 

Summary of results: JOSO"C Observed viscosity (cp) 1.787 0.2821 Predicked by Eq. 1.5-10 2.57 Production by Eq. 1.5-12 0.95

Both equations give poor predictions. This is not surprising inasmuch as the empirical formula in Eqs. 1.5-7 stacy. do not hold for wake, nor for most associated sub-Sittematics.

## 1. E Molecular Velocity and Mean Free Park

From Eq. 1.4-1 the mean speed of a molecule is

 $\overline{u} = \sqrt{\frac{8RT}{\pi M}} = \sqrt{\frac{8(8.314 \times 10^7)(273.2)}{\pi(32.00)}} = 4.25 \times 10^4 \text{ cm sec}^4$ 

From Eq. 1.4-3 the mean free path is: 
$$\lambda = \frac{RT}{\sqrt{2} \pi d^2 p N} = \frac{(82.057)(273.2)}{\sqrt{2} \pi (3 \times 10^{-8})^2 (1)(6.023 \times 10^{23})} = 9.3 \times 10^{-6}$$
 cm

Hence the mean free path is  $(9.3 \times 10^6)/(3 \times 10^{-3}) = 310$  molecular diumetro under these conditions. In the liquid state, on the other hand, the corresponding ratio would be of the order of magnitude of or even less than, unity.

# 1.F Comparison of the Uyehara-Watson Chart with Kinetic Theory

(a) Combination of Eqs. 1.4-11 and 18 gives.

In addition, from Eq. 1.4-11 we know that T = 0.77 &T/E The calculations for the kinetic theory plot are summarized below:

# **Solutions Manual Transport Phenomena And Unit Operations**

**Tao Wei** 

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