

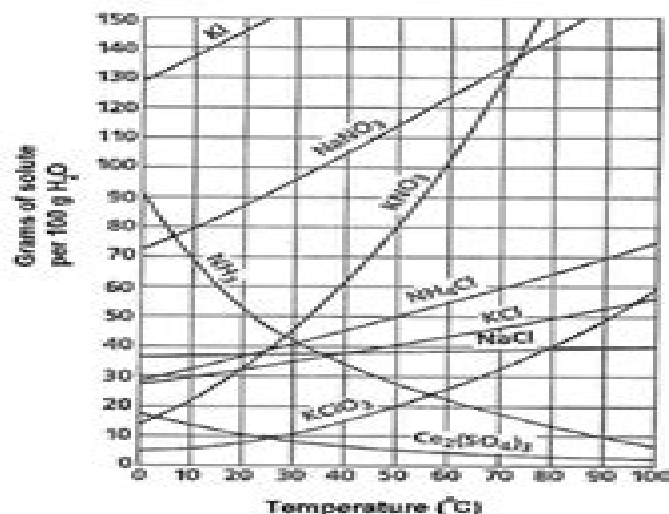
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Solutions Study Guide



1. Using the solubility curves above, answer the following questions:

- a. Which substance will form a saturated solution when 35 g is dissolved in 100 g of water at 75°C? Answer in a complete sentence.

Potassium chlorate (KClO_3) will form a saturated solution when 35g are dissolved in 100g of water at 75°C.

- b. Which substance(s) on the graph are gases? Justify your response using complete sentences.

NH_3 is a gas. The solubility of gases decrease as the temperature increases. Since the slope of the curve for NH_3 is negative, this indicates a decrease in solubility with increased temperature.

- c. Which substance is LEAST affected by temperature? Justify your response using complete sentences.

Sodium chloride (NaCl) is least affected by temperature. This is indicated in the small slope of its curve.

- d. If 65 g of ammonium chloride is dissolved in 100 g of water at 90°C, will the solution be saturated, unsaturated or supersaturated. Justify your response using complete sentences.

The solution will be unsaturated. The point for 65g of NH_4Cl at 90°C falls below the curve indicating that more solute could be dissolved.

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Ben Piper, David Clinton



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$a^2 + b^2 = c^2 \cos^2 \theta$

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