Chapter 2 Diffusion in dilute solutions

1. Water evaporation

To find the rate of evaporation, we need to find the flux of water across the air film:

$$\dot{f_1} = \frac{DH}{\ell} \Delta c_1$$

Since the film is made up of air, H = 1. We are given D and ℓ , so we need to find de_{ℓ} . We first calculate the difference in the partial pressure of water across the film. From a steam table we find that at 20 °C, $p_{tot} = 2.3388$ kPa. Since the air immediately above the water is saturated we have:

$$\Delta p_i = p_{im} - 0.5 p_{im} = 0.5 p_{im} = 0.5(2.3388 kPa) = 1.1694 kPa = 1169.4 Pa$$

Assuming an ideal gas, we can find the concentration difference:

$$\Delta c_1 = \frac{\Delta p_1}{RT} = \frac{1169.4 \, Pa}{\left(8.3145 \, \frac{J}{K \cdot mol}\right) (293 \, K)} = 0.48 \, \frac{mol}{m^3} = 4.8 \cdot 10^{-7} \, \frac{mol}{cm^3}$$

We can now calculate the flux:

$$J_1 = \frac{0.25 \frac{cm^2}{s}}{0.15 cm} \left(4.8 \cdot 10^{-7} \frac{mol}{cm^3} \right) = 8.0 \cdot 10^{-7} \frac{mol}{cm^2 s}$$

To find the height change, we divide by the molar density of water:

$$\frac{8.0 \cdot 10^{-3} \frac{mol}{cm^2 s}}{1 \frac{g}{cm^3} \cdot \frac{1 mol}{18.015 g}} = 1.44 \cdot 10^{-3} \frac{cm}{s} = 1.25 \frac{cm}{day}$$

2. Diffusion across a monolayer

Recalling that the resistance is the inverse of the permeance, we have:

$$\frac{1}{P} = 2 \frac{s}{cm} \Rightarrow P = 0.5 \frac{cm}{s}$$

We know that the permeance is given by:

$$P = \frac{DH}{\ell} \Rightarrow D = \frac{P\ell}{H} = \frac{\left(0.5 \frac{cm}{s}\right) \left(2.5 \cdot 10^{-7} cm\right)}{0.018} = 7 \cdot 10^{-6} \frac{cm^2}{s}$$

3. Diffusion coefficient of NO; in water

We treat the water as a semiinfinite slab. From Eq. (2.3-18) we know that the flux at the surface is given by:

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Table of Contents Solutions Manual To Cussler 3rd Edition

- 1. Understanding the eBook Solutions Manual To Cussler 3rd Edition
 - The Rise of Digital Reading Solutions Manual To Cussler 3rd Edition
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Solutions Manual To Cussler 3rd Edition
 - Exploring Different Genres
 - o Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Solutions Manual To Cussler 3rd Edition
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Solutions Manual To Cussler 3rd Edition

- Personalized Recommendations
- Solutions Manual To Cussler 3rd Edition User Reviews and Ratings
- Solutions Manual To Cussler 3rd Edition and Bestseller Lists
- 5. Accessing Solutions Manual To Cussler 3rd Edition Free and Paid eBooks
 - Solutions Manual To Cussler 3rd Edition Public Domain eBooks
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 - Interactive Elements Solutions Manual To Cussler 3rd Edition
- 8. Staying Engaged with Solutions Manual To Cussler 3rd Edition
 - o Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Solutions Manual To Cussler 3rd Edition
- 9. Balancing eBooks and Physical Books Solutions Manual To Cussler 3rd Edition
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Solutions Manual To Cussler 3rd Edition
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Solutions Manual To Cussler 3rd Edition
 - Setting Reading Goals Solutions Manual To Cussler 3rd Edition
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Solutions Manual To Cussler 3rd Edition

- Fact-Checking eBook Content of Solutions Manual To Cussler 3rd Edition
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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