

### **Worksheet #3—Stoichiometry**

**Directions:** Solve the following stoichiometry problems. Assume all gases are at STP. Don't forget to write out your balanced chemical equation. Make sure to show ALL work to receive FULL CREDIT!

1. Calculate the number of moles of oxygen required to react completely with 14.8 mol Al in the following equation:  $4\text{Al}_{(s)} + 3\text{O}_{2(g)} \rightarrow 2\text{Al}_2\text{O}_{3(s)}$

$$\frac{14.8 \text{ moles Al}}{4 \text{ mole Al}} \times \frac{3 \text{ moles of O}_2}{1 \text{ mole Al}} = \underline{11.1 \text{ mol O}_2}$$

3. If 20.0g grams of zinc react with excess hydrochloric acid, how many grams of zinc chloride ( $\text{ZnCl}_2$ ) are produced?



$$\frac{20 \text{ g Zn}}{65 \text{ g Zn}} \times \frac{1 \text{ mol Zn}}{1 \text{ mole Zn}} \times \frac{1 \text{ mole ZnCl}_2}{1 \text{ mole Zn}} \times \frac{136 \text{ g ZnCl}_2}{1 \text{ mol ZnCl}_2}$$

$$= \underline{41.8 \text{ g ZnCl}_2}$$

5. How many grams of copper are required to replace 4.00 grams of silver nitrate which are dissolved in water?



$$\frac{4.0 \text{ g AgNO}_3}{170 \text{ g AgNO}_3} \times \frac{1 \text{ mol AgNO}_3}{1 \text{ mole AgNO}_3} \times \frac{2 \text{ mole AgNO}_3}{1 \text{ mole Cu}} \times \frac{63.5 \text{ g Cu}}{1 \text{ mole Cu}}$$

2. Calculate the moles of oxygen produced from the decomposition of 0.61 moles of potassium chlorate ( $\text{KClO}_3$ ).



$$\frac{0.61 \text{ moles KClO}_3}{2 \text{ mole KClO}_3} \times \frac{3 \text{ moles of O}_2}{1 \text{ mole KClO}_3} =$$

$$\underline{0.915 \text{ mol O}_2}$$

4. How many grams of chlorine gas must be reacted with excess sodium iodide (NaI) if 10.0 grams of sodium chloride (NaCl) are needed



$$\frac{10 \text{ g NaCl}}{58.5 \text{ g NaCl}} \times \frac{1 \text{ mol NaCl}}{1 \text{ mole NaCl}} \times \frac{1 \text{ mole Cl}_2}{2 \text{ mole NaCl}} \times \frac{71 \text{ g Cl}_2}{1 \text{ mole Cl}_2}$$

$$= \underline{6.07 \text{ g Cl}_2}$$

6. How many moles of hydrogen gas are needed to react with 15.1 g chlorine gas to produce hydrogen chloride gas?



$$\frac{15.1 \text{ g Cl}_2}{71 \text{ g Cl}_2} \times \frac{1 \text{ mol Cl}_2}{1 \text{ mole Cl}_2} \times \frac{1 \text{ mole H}_2}{1 \text{ mole Cl}_2}$$

$$= \underline{0.21 \text{ moles H}_2}$$

# Solving Problems A Chemistry H Answer Key

**Marjorie Gardner, James G.  
Greeno, Frederick Reif, Alan H.  
Schoenfeld, Andrea A. diSessa**



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**A Problem-Solving Approach to Aquatic Chemistry** James N. Jensen, 2023-01-05 A Problem Solving Approach to Aquatic Chemistry Enables civil and environmental engineers to understand the theory and application of aquatic equilibrium chemistry The second edition of A Problem Solving Approach to Aquatic Chemistry provides a detailed introduction to aquatic equilibrium chemistry calculation methods for systems at equilibrium applications of aquatic chemistry and chemical kinetics The text directly addresses two required ABET program outcomes in environmental engineering chemistry including stoichiometry equilibrium and kinetics and material and energy balances fate and transport of substances in and between air water and soil phases The book is very student centered with each chapter beginning with an introduction and ending with a summary that reviews the chapter's main points To aid in reader comprehension important terms are defined in context and key ideas are summarized Many thought provoking discussion questions worked examples and end of chapter problems are also included Each part of the text begins with a case study a portion of which is addressed in each subsequent chapter illustrating the principles of that chapter In addition each chapter has an Historical Note exploring connections with the people and cultures connected to topics in the text A Problem Solving Approach to Aquatic Chemistry includes Fundamental concepts such as concentration units thermodynamic basis of equilibrium and manipulating equilibria Solutions of chemical equilibrium problems including setting up the problems and algebraic graphical and computer solution techniques Acid base equilibria including the concepts of acids and bases titrations and alkalinity and acidity Complexation including metals ligands equilibrium calculations with complexes and applications of complexation chemistry Oxidation reduction equilibria including equilibrium calculations graphical approaches and applications Gas liquid and solid liquid equilibrium with expanded coverage of the effects of global climate change Other topics including chemical kinetics of aquatic systems surface chemistry and integrative case studies For advanced senior undergraduates and first year graduate students in environmental engineering courses A Problem Solving Approach to Aquatic Chemistry serves as an invaluable learning resource on the topic with a variety of helpful learning elements included throughout to ensure information retention and the ability to apply covered concepts in practical settings

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**Chemistry: The Central Science**  
 Theodore L. Brown, H. Eugene LeMay Jr., Bruce E. Bursten, Catherine Murphy, Patrick Woodward, Steven Langford, Dalius Sagatys, Adrian George, 2013-10-04 If you think you know the Brown LeMay Bursten Chemistry text think again In response to market request we have created the third Australian edition of the US bestseller Chemistry The Central Science An extensive revision has taken this text to new heights Triple checked for scientific accuracy and consistency this edition is a more seamless and cohesive product yet retains the clarity innovative pedagogy functional problem solving and visuals of the previous version All artwork and images are now consistent in quality across the entire text And with a more traditional and logical organisation of the Organic Chemistry content this comprehensive text is the source of all the information and practice problems students are likely to need for conceptual understanding development of problem solving skills reference and test preparation

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Stereochemistry & Mechanism Through solved Problems , **CAT | COMMON ADMISSION TEST| 10 YEARS'| CHAPTERWISE & TOPICWISE| SOLVED PAPERS|QUESTION BANK|2006- 2008|2017 - 2023 (Subject Name - VARC, DILR & QA)** Oswaal Editorial Board, 2024-09-05 Other CAT Books Common Admission Test CAT Common Admission Test CAT is the gateway to all the prestigious management colleges in India including the coveted IIMs Indian Institutes of Management It is a distinctive exam in the context that it assesses a candidate's logical ability endurance and presence of mind thus it cannot be qualified by just rote learning So as long as the aspirants are ready to work hard and sharpen their minds they can clear CAT with flying colours To ensure that the aspirants names appear in the merit lists of top colleges it is necessary to pick the right study material and conduct their preparation strategically Oswaal CAT Chapter wise Topic wise consists of Previous 10 Years Solved Papers 2006 2008 and 2017 2023 prepared by Oswaal Editorial Board after thorough research and analysis of the Exam Pattern Syllabus that has been followed year on year We are giving the actual years of questions wherein 2009 2016 papers are not shared by the IIMs This book has just the right ingredients to help the aspirants crack CAT 2024 Like every year this year also the exam is expected to be conducted on the last Sunday of November 2024 Key Benefits 100% Updated With 2023 Paper Fully Solved Extensive Practice With 1200 Questions Detailed Explanations Concept Clarity learn key

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Benjamin, 2014-09-26 Aquatic chemistry students need a solid foundation in fundamental concepts as well as numerical techniques for solving the variety of problems they will encounter as practicing engineers For over a decade Mark Benjamin's Water Chemistry has brought to the classroom a balanced coverage of fundamentals and analytical algorithms in a student friendly accessible way The text distinguishes itself with longer and more detailed explanations of the relevant chemistry and mathematics allowing students to understand not only which techniques work best for a given application but also why those techniques should be applied and what their limitations are The end result is a solid thorough framework for comprehending equilibrium in complex aquatic systems The second edition includes a thorough introductory explanation of chemical reactivity and a new chapter on reaction kinetics providing much needed context as well as full treatments of the tableau method and TOTH equation The discussion of the thermodynamic perspective on chemical reactivity has been extensively revised The entire book now integrates Visual Minteq the most popular software for analyzing chemical equilibria into the problem solving approach Additional exercises range more widely in difficulty giving instructors more flexibility and diversity in their assignments **Organic Chemistry** T. W. Graham Solomons, Craig B. Fryhle, Scott A. Snyder, 2016-01-19 The 12th

edition of Organic Chemistry continues Solomons Fryhle Snyder's tradition of excellence in teaching and preparing students for success in the organic classroom and beyond A central theme of the authors approach to organic chemistry is to emphasize the relationship between structure and reactivity To accomplish this the content is organized in a way that combines the most useful features of a functional group approach with one largely based on reaction mechanisms The authors philosophy is to emphasize mechanisms and their common aspects as often as possible and at the same time use the unifying features of functional groups as the basis for most chapters The structural aspects of the authors approach show students what organic chemistry is Mechanistic aspects of their approach show students how it works And wherever an opportunity arises the authors show students what it does in living systems and the physical world around us **Catalog of**

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[Introduction to General, Organic, and Biochemistry](#) Morris Hein, Scott Pattison, Susan Arena, Leo R. Best, 2014-01-15 The most comprehensive book available on the subject Introduction to General Organic and Biochemistry 11th Edition continues its tradition of fostering the development of problem solving skills featuring numerous examples and coverage of current applications Skillfully anticipating areas of difficulty and pacing the material accordingly this readable work provides clear and logical explanations of chemical concepts as well as the right mix of general chemistry organic

chemistry and biochemistry An emphasis on real world topics lets readers clearly see how the chemistry will apply to their career

Ebook: Organic Chemistry Janice Smith, 2014-10-16 Serious Science with an Approach Built for Today's Students Smith's Organic Chemistry continues to breathe new life into the organic chemistry world This new fourth edition retains its popular delivery of organic chemistry content in a student friendly format Janice Smith draws on her extensive teaching background to deliver organic chemistry in a way in which students learn with limited use of text paragraphs and through concisely written bulleted lists and highly detailed well labeled teaching illustrations Don't make your text decision without seeing Organic Chemistry 4th edition by Janice Gorzynski Smith

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**Toward a Scientific Practice of Science Education** Marjorie Gardner, James G. Greeno, Frederick Reif, Alan H. Schoenfeld, Andrea A. diSessa, 2013-04-03 This volume supports the belief that a revised and advanced science education can emerge from the convergence and synthesis of several current scientific and technological activities including examples of research from cognitive science social science and other discipline based educational studies The anticipated result the formation of science education as an integrated discipline

**Analytical Chemistry, International Adaptation** Gary D. Christian, Purnendu K. Dasgupta, Kevin A. Schug, 2020 With the 7th Edition of Analytical Chemistry renowned chemists Purnendu Sandy Dasgupta and Kevin Schug both of the University of Texas Arlington join the author team The new edition focuses on more in depth coverage of the principles and techniques of quantitative analysis and instrumental analysis aka Analytical Chemistry The goal of the text is to provide a foundation of the analytical process tools and computational methods and resources and to illustrate with problems that bring realism to the practice and importance of analytical chemistry It is designed for undergraduate college students majoring in chemistry and in fields related to chemistry

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