



COVALENT BONDING

Practice Problems

Do your answers to these very practice problems.

SECTION 8.1 MOLECULAR COMPOUNDS

1. Classify each of the following compounds as ionic or covalent.

- a. Na_2O (ionic) b. CH_4 (covalent) c. HCl (covalent)
 d. KNO_3 (ionic) e. H_2O (covalent)

2. Write the formula for the following molecular compounds.

- a. carbon dioxide b. sulfur dioxide c. carbon monoxide
 d. N_2O e. SO_2 f. CO

3. Write the Lewis structure for each of the following molecular compounds.

a. CH_4 (carbon has 4 valence electrons and hydrogen has 1; there are 10 valence electrons in total and 4 bonds must be formed in the molecule)

b. H_2O (oxygen has 6 valence electrons and hydrogen has 1; there are 8 valence electrons in total and 2 bonds must be formed in the molecule)

SECTION 8.2 THE NATURE OF COVALENT BONDING

1. Draw the Lewis structure for hydrogen fluoride, HF .



2. Draw the Lewis structure for methane, CH_4 .



3. Draw the Lewis structure for carbon dioxide, CO_2 .

4. Draw the Lewis structure for acetylene, C_2H_2 .



5. Draw the Lewis structure for ammonia, NH_3 . Be sure to show all lone pairs of electrons.



(Lone pairs are shown as pairs of dots.)

SECTION 8.3 BONDING: THE ORBITAL

1. Which orbital has two regions of electron density, one above and one below the nucleus?

2. Which orbital has two regions of electron density, one to the left and one to the right of the nucleus?

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Barbara Kirchner



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Journal of the Indian Chemical Society Indian Chemical Society,1983

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