

## Answers

- 55  $M = 133 \text{ g mol}^{-1}$  so gas is Xe
- 56  $90.4 \text{ g mol}^{-1}$
- 57 Helium
- 58  $311 \text{ dm}^3$
- 59 empirical formula and molecular formula =  $\text{SO}_2$
- 60 At higher altitude the external pressure is less. As the air in the tyre expands on heating (due to friction with the road surface), the internal pressure increases.
- 61 (a) Particles are in constant random motion and collide with each other and with the walls of the container in perfectly elastic collisions. The kinetic energy of the particles increases with temperature. There are no inter-particle forces and the volume of the particles is negligible relative to the volume of the gas.
- (b) At low temperature, the particles have lower kinetic energy, which favours the formation of inter-particle forces and reduces gas pressure.  $\frac{PV}{nRT} < 1$
- 62  $\text{NH}_3$  shows greater deviation than  $\text{CH}_4$  due to stronger intermolecular attractions, especially at low temperature.
- 63 B
- 64  $2.81 \text{ g}$
- 65  $4.93 \text{ g}$
- 66  $0.0100 \text{ mol}$
- 67  $0.400 \text{ mol dm}^{-3}$
- 68  $3.1 \text{ cm}^3$
- 69  $0.178 \text{ mol dm}^{-3}$
- 70  $0.0220 \text{ mol dm}^{-3}$ ,  $0.0602\% \text{ HCl}$
- 71  $0.106 \text{ mol dm}^{-3} \text{ Na}_2\text{SO}_4$  and  $0.115 \text{ mol dm}^{-3} \text{ Pb}(\text{NO}_3)_2$ ; assume no side reactions, all  $\text{PbSO}_4$  precipitates
- 72 1217 tonnes

4

- 73  $52\% \text{ NH}_3$  by mass; assuming no side reactions occur and gases behave as ideal gases
- 74  $3.20 \times 10^2 \text{ kg}$
- 75  $0.225 \text{ mol dm}^{-3}$  (or round to  $0.23 \text{ mol dm}^{-3}$ )
- 76  $[\text{Na}_2\text{CO}_3] = \frac{VP}{x} \text{ mol dm}^{-3}$   
 $[\text{NaHCO}_3] = \frac{V(P - P)}{x} \text{ mol dm}^{-3}$
- 77  $100 \text{ mol O}_2$ ;  $3.2 \text{ kg}$

## Challenge yourself

- In cold climates, temperature may approach or go below the boiling point of butane so the butane stays liquid even when it is released from the pressure it is under when stored in its canister. This makes it ineffective as a fuel.
- $\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$ ,  $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ ,  $\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$
- N = 16%, P = 22%, K = 17%
- Many reactions with "useless" by-products could have high stoichiometric yield under optimum conditions, but low atom economy, for example, methanoic acid production:  
 $2\text{NaCOOH} + \text{H}_2\text{SO}_4 \rightarrow 2\text{HCOOH} + \text{Na}_2\text{SO}_4$   
 For 100% conversion with stoichiometric reactants, the yield = 100%.  

$$\text{atom economy} = \frac{2 \times 46.03}{(2 \times 68.01) + 98.08} \times 100\% = 39\%$$
- $2\text{NaN}_3(\text{s}) \rightarrow 2\text{Na}(\text{s}) + 3\text{N}_2(\text{g})$   
 $10\text{Na}(\text{s}) + 2\text{KNO}_3(\text{s}) \rightarrow \text{K}_2\text{O}(\text{s}) + 5\text{Na}_2\text{O}(\text{s}) + \text{N}_2(\text{g})$   
 $\text{K}_2\text{O}(\text{s}) + \text{Na}_2\text{O}(\text{s}) + \text{SiO}_2(\text{s}) \rightarrow \text{Na}_2\text{K}_2\text{SiO}_5$  (alkaline silicate glass)
- As  $\text{NaOH}$  dissolves, the separated  $\text{Na}^+$  and  $\text{OH}^-$  ions become hydrated, i.e. they are surrounded by  $\text{H}_2\text{O}$  molecules. This involves breaking the hydrogen bonds between the  $\text{H}_2\text{O}$  molecules in pure water and allows closer packing, which reduces the volume.

# Pearson Chemistry Chapter 22 Workbook Answers

**David Kirk**



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### **Table of Contents Pearson Chemistry Chapter 22 Workbook Answers**

1. Understanding the eBook Pearson Chemistry Chapter 22 Workbook Answers
  - The Rise of Digital Reading Pearson Chemistry Chapter 22 Workbook Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Pearson Chemistry Chapter 22 Workbook Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Pearson Chemistry Chapter 22 Workbook Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Pearson Chemistry Chapter 22 Workbook Answers

- Personalized Recommendations
  - Pearson Chemistry Chapter 22 Workbook Answers User Reviews and Ratings
  - Pearson Chemistry Chapter 22 Workbook Answers and Bestseller Lists
5. Accessing Pearson Chemistry Chapter 22 Workbook Answers Free and Paid eBooks
- Pearson Chemistry Chapter 22 Workbook Answers Public Domain eBooks
  - Pearson Chemistry Chapter 22 Workbook Answers eBook Subscription Services
  - Pearson Chemistry Chapter 22 Workbook Answers Budget-Friendly Options
6. Navigating Pearson Chemistry Chapter 22 Workbook Answers eBook Formats
- ePub, PDF, MOBI, and More
  - Pearson Chemistry Chapter 22 Workbook Answers Compatibility with Devices
  - Pearson Chemistry Chapter 22 Workbook Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Pearson Chemistry Chapter 22 Workbook Answers
  - Highlighting and Note-Taking Pearson Chemistry Chapter 22 Workbook Answers
  - Interactive Elements Pearson Chemistry Chapter 22 Workbook Answers
8. Staying Engaged with Pearson Chemistry Chapter 22 Workbook Answers
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Pearson Chemistry Chapter 22 Workbook Answers
9. Balancing eBooks and Physical Books Pearson Chemistry Chapter 22 Workbook Answers
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Pearson Chemistry Chapter 22 Workbook Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Pearson Chemistry Chapter 22 Workbook Answers
- Setting Reading Goals Pearson Chemistry Chapter 22 Workbook Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Pearson Chemistry Chapter 22 Workbook Answers

- Fact-Checking eBook Content of Pearson Chemistry Chapter 22 Workbook Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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