

## Reviewing Content

42. The solvent is the substance in which the solute is dissolved.
43. Random collisions of the solvent molecules with the solute particles provide enough force to overcome gravity.
44. **solubility**: the amount of a substance that dissolves in a given quantity of solvent at specified conditions of temperature and pressure to produce a saturated solution.  
**saturated solution**: a solution containing the maximum amount of solute for a given amount of solvent at a constant temperature and pressure. **unsaturated solution**: a solution that contains less solute than a saturated solution at a given temperature and pressure. **miscible**: describes liquids that dissolve in each other. **immiscible**: describes liquids that are insoluble in each other.
45. Particles of solute crystallize.
46. No; if there were undissolved solute, the excess solute would come out of a supersaturated solution.
47.  $5.55 \times 10^2$  g  $\text{AgNO}_3$
48. Solubility increases with pressure.
49. a.  $1.6 \times 10^{-2}$  g/L  
b.  $4.7 \times 10^{-2}$  g/L
50. *Dilute* and *concentrated* are relative terms and are not quantitative. Molarity provides the exact number of moles of solute per liter of solution.
51. Molarity is the number of moles of solute dissolved in one liter of solution.  
a. 1.3M KCl  
b.  $3.3 \times 10^{-1}$ M  $\text{MgCl}_2$
52.  $2.00 \times 10^3$  mL
53. a.  $5.0 \times 10^{-3}$  mol NaCl, 29 g NaCl  
b. 1.0 mol  $\text{KNO}_3$ ,  $1.0 \times 10^{-2}$  g  $\text{KNO}_3$   
c.  $2.5 \times 10^3$  mol  $\text{CaCl}_2$ , 2.8 g  $\text{CaCl}_2$
54. a.  $2.3 \times 10^1$  g NaCl  
b. 2.0 g  $\text{MgCl}_2$
55. a. 16% (v/v) ethanol  
b. 63.6% (v/v) isopropyl alcohol
56. Colligative properties are properties of a solution that depend only on the number of solute particles; boiling-point elevation, freezing-point depression, and vapor-pressure lowering. Boiling points are elevated because shells of solvent form around solute particles, reducing the amount of solvent molecules that have sufficient energy to escape the solution; relative to the pure solvent, the amount of energy required to cause vaporization or boiling increases. Solutes disrupt the ordering of the solvent structure, so more kinetic energy must be withdrawn from a solution for it to solidify. This lowers the freezing point of the solution.
57. a. sea water  
b. 1.5M  $\text{KNO}_3$   
c. 0.100M  $\text{MgCl}_2$
58. The effective molality of the  $\text{Ca}(\text{NO}_3)_2$  solution is 3m. The effective molality of the  $\text{NaNO}_3$  solution is 2m.
59. When vapor pressure is lowered relative to pure solvent, more energy must be supplied to reach the boiling point; thus the boiling point is increased relative to pure solvent.
60. The salt lowers the freezing point of the ice-water cooling mixture.
61. 1M solution: 1 mol of solute in 1 L of solution; 1m solution: 1 mol of solute in 1000 g of solvent
62. Add 27.0 g  $\text{H}_2\text{O}$  to 32.0 g  $\text{CH}_3\text{OH}$ .
63. a.  $100.26^\circ\text{C}$   
b.  $101.54^\circ\text{C}$
64. a.  $-4.46^\circ\text{C}$   
b.  $-2.2^\circ\text{C}$
65. a.  $-1.1^\circ\text{C}$   
b.  $-0.74^\circ\text{C}$   
c.  $-1.5^\circ\text{C}$

## Understanding Concepts

66. a. The freezing-point depression is twice as great for solute B; solute B must provide twice as many particles in solution.

# Pearson Education Chemistry Packet Answers 16

**Albert A Gayle**



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### **Table of Contents Pearson Education Chemistry Packet Answers 16**

1. Understanding the eBook Pearson Education Chemistry Packet Answers 16
  - The Rise of Digital Reading Pearson Education Chemistry Packet Answers 16
  - Advantages of eBooks Over Traditional Books
2. Identifying Pearson Education Chemistry Packet Answers 16
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an eBook Pearson Education Chemistry Packet Answers 16
  - User-Friendly Interface
4. Exploring eBook Recommendations from Pearson Education Chemistry Packet Answers 16
  - Personalized Recommendations
  - Pearson Education Chemistry Packet Answers 16 User Reviews and Ratings
  - Pearson Education Chemistry Packet Answers 16 and Bestseller Lists

5. Accessing Pearson Education Chemistry Packet Answers 16 Free and Paid eBooks
  - Pearson Education Chemistry Packet Answers 16 Public Domain eBooks
  - Pearson Education Chemistry Packet Answers 16 eBook Subscription Services
  - Pearson Education Chemistry Packet Answers 16 Budget-Friendly Options
6. Navigating Pearson Education Chemistry Packet Answers 16 eBook Formats
  - ePub, PDF, MOBI, and More
  - Pearson Education Chemistry Packet Answers 16 Compatibility with Devices
  - Pearson Education Chemistry Packet Answers 16 Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Pearson Education Chemistry Packet Answers 16
  - Highlighting and Note-Taking Pearson Education Chemistry Packet Answers 16
  - Interactive Elements Pearson Education Chemistry Packet Answers 16
8. Staying Engaged with Pearson Education Chemistry Packet Answers 16
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Pearson Education Chemistry Packet Answers 16
9. Balancing eBooks and Physical Books Pearson Education Chemistry Packet Answers 16
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Pearson Education Chemistry Packet Answers 16
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Pearson Education Chemistry Packet Answers 16
  - Setting Reading Goals Pearson Education Chemistry Packet Answers 16
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Pearson Education Chemistry Packet Answers 16
  - Fact-Checking eBook Content of Pearson Education Chemistry Packet Answers 16
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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