#### Part B True-False

8.	AT	IO. ST	12. NT
9.	PATE	11. NT	

#### Part C Matching

13. a	15. b	17. d
1.4 0	16 0	

#### Part D Questions and Problems

20. 
$$\Delta H = 5.53 \frac{\text{mol NH}_4\text{NO}_2\text{UU}}{25.7 \text{ kJ}} \times \frac{25.7 \text{ kJ}}{1 \frac{\text{mol NH}_5\text{NO}_2\text{UU}}{1.000 \text{ kJ}_4\text{NO}_2\text{UU}}} = 142 \text{ kJ}$$

#### Section 17.4

#### Part A Completion

- sum
   enthalpy
- 2. enthalpy
- 3. indirectly
- 4. changed (reversed)
- 5. standard heat of formation
- 6. change
- 7. come
- 8. 4//
- 9. pero
- 10. subtracting

#### Part B True-False

11. NT 13. AT 15. NT 12. NT 14. NT

#### Part C Matching

16. b 18. a 20. c 17. c 19. d

#### Part D Questions and Problems

21. 
$$CuO(s) \rightarrow Cu(s) + \frac{1}{2}O_2 \Delta H = +155 \text{ kJ}$$
  
 $H_2(g) + \frac{1}{2}O_2(g) \rightarrow +H_2O(g) \Delta H = -242 \text{ kJ}$   
 $CuO(s) + H_2(g) \rightarrow Cu(s) + H_2O(g) \Delta H_{min} = -87 \text{ kJ}$ 

22. 
$$\Delta H_i^0$$
 (products)  
=  $3 \mod \Theta_i^0$  ×  $\frac{-393.5 \text{ kJ}}{1 \mod \Theta_i^0}$ ;  
=  $-1181 \text{ kJ}$   
 $\Delta H_i^0$  (reactants)  
=  $3 \mod \Theta_i^0$  ×  $\frac{-110.5 \text{ kJ}}{1 \mod \Theta_i^0}$  +  $(-822.1 \text{ kJ})$   
=  $-1154 \text{ kJ}$   
 $\Delta H^0 = \Delta H_i^0$  (products)  $-\Delta H_i^0$  (reactants)  
=  $-1181 \text{ kJ} - (-1154 \text{ kJ})$   
=  $-27 \text{ kJ}$ 

#### **Practice Problems**

#### Section 17.1

1. 
$$200.0 \text{ Cal} \times \frac{1000 \text{ cal}}{1 \text{ Cal}} \times \frac{4.184 \text{ J}}{1 \text{ cal}} \times \frac{1 \text{ kJ}}{10^3 \text{ J}}$$
  
 $= 836.8 \text{ kJ}$   
2.  $C = \frac{525.0 \text{ cal}}{25.0 \text{ g} \times 15.0 ^{\circ}\text{C}} = 1.40 \text{ cal/(g} \times ^{\circ}\text{C)}$   
3.  $\Delta T = \frac{1255.0 \text{ J}}{100.0 \text{ g} \times 2.1 \text{ J/(g} \times ^{\circ}\text{C)}} = 6.0 ^{\circ}\text{C}$   
4.  $q = 100.0 \text{ g} \times 120.0 ^{\circ}\text{C} \times 0.90 \frac{\text{J}}{\text{g} \times ^{\circ}\text{C}}$   
 $= 1.1 \times 10^3 \text{ J}$ 

#### Section 17.2

1. 
$$\Delta H = 150.0 \text{ g} \times 4.18 \frac{\text{J}}{\text{g} \times \text{°C}} \times 10^{\circ}\text{C}$$
  
=  $6.3 \times 10^3 \text{ J} = 6.3 \text{ kJ}$ 

2. 
$$\Delta H = 15.0 \text{ g CasOH5}_2(3) \times \frac{1 \text{ mol CasOH5}_2(3)}{74.1 \text{ g CasOH5}_2(3)} \times \frac{-65.2 \text{ kJ}}{1 \text{ mol CasOH5}_2(3)} = -13.2 \text{ kJ}$$

3. 
$$\Delta H = 52.4 \text{ g.CH2(S)} \times \frac{1 \text{ mol GH2(S)}}{16.0 \text{ g.CH2(S)}} \times \frac{-890.2 \text{ kJ}}{1 \text{ mol GH2(S)}} = 2.93 \times 10^3 \text{ kJ}$$

4. 
$$4NH_3 + 5O_2(g) \rightarrow 4NO(g) + 6H_2O(g)$$
  

$$\Delta H = \frac{-226 \text{ kJ}}{1 \text{ mol NH}_2(g)} \times 4 \text{ mol NH}_2(g)$$

$$= -904 \text{ kJ}$$

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Judd E. Hollander

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