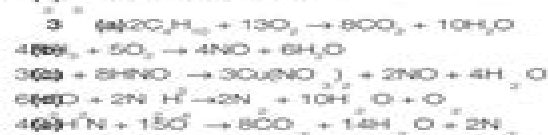
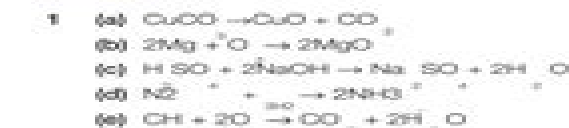


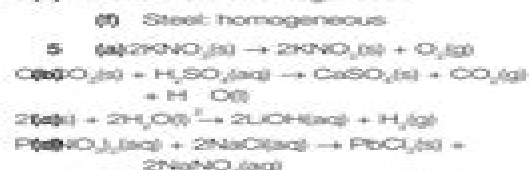
Answers

Chapter 1

Exercises



4 (a) Sand and water: heterogeneous
 (b) Sugar and water: homogeneous
 (c) Oil and iron filings: heterogeneous
 (d) Endol and water: homogeneous



In the answers, exercises and text there is inconsistency about whether the water formed in combustion reactions is gas or liquid. At room temperature it will ultimately finish up as a liquid, as shown here.

6 X has diffused more quickly, so it must be a lighter gas. Its particles have greater velocity than the particles of Y at the same temperature. (Note though that they will both have the same value for average kinetic energy.)

7 From the kinetic molecular theory we would expect a solid to be more dense than its liquid, and therefore that ice would sink in water.

8 Bubbles will be present through the volume of the liquid. A brown gas is visible above the brown liquid. As the two states are at the same temperature, the particles have the same average kinetic energy and are moving at the same speed. The inter-particle distances in the gas are significantly larger than those in the liquid.

9 At certain conditions of low temperature and low humidity, snow changes directly to water vapour by sublimation, without going through the liquid phase.

10 Steam will condense on the skin, releasing energy as it forms liquid at the same temperature (E–D on Figure 1.4). This is additional to the energy released when both the boiling water and the condensed steam cool on the surface of the skin.

11 B

12 AW 01_EX_12

Pearson Chemistry 3 Answer Key

Ananya Ganguly



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1. Understanding the eBook Pearson Chemistry 3 Answer Key
 - The Rise of Digital Reading Pearson Chemistry 3 Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Pearson Chemistry 3 Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pearson Chemistry 3 Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pearson Chemistry 3 Answer Key
 - Personalized Recommendations
 - Pearson Chemistry 3 Answer Key User Reviews and Ratings

- Pearson Chemistry 3 Answer Key and Bestseller Lists
- 5. Accessing Pearson Chemistry 3 Answer Key Free and Paid eBooks
 - Pearson Chemistry 3 Answer Key Public Domain eBooks
 - Pearson Chemistry 3 Answer Key eBook Subscription Services
 - Pearson Chemistry 3 Answer Key Budget-Friendly Options
- 6. Navigating Pearson Chemistry 3 Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Pearson Chemistry 3 Answer Key Compatibility with Devices
 - Pearson Chemistry 3 Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pearson Chemistry 3 Answer Key
 - Highlighting and Note-Taking Pearson Chemistry 3 Answer Key
 - Interactive Elements Pearson Chemistry 3 Answer Key
- 8. Staying Engaged with Pearson Chemistry 3 Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pearson Chemistry 3 Answer Key
- 9. Balancing eBooks and Physical Books Pearson Chemistry 3 Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pearson Chemistry 3 Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Pearson Chemistry 3 Answer Key
 - Setting Reading Goals Pearson Chemistry 3 Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pearson Chemistry 3 Answer Key
 - Fact-Checking eBook Content of Pearson Chemistry 3 Answer Key
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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