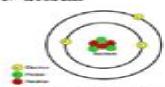
#### Atomic Structure – Study Guide Grade S – Physical Science

Atom -- smallest particle of an element that has all the properties of that element.

#### 3 main subatomic particles:

- I. protons
- 2. neutrons
- 3. electrons



nucleus -- core of the atom, it is positively charged 99.9 % of the atoms mass is in the nucleus.

Protons and neutrons are found in the nucleus.

Protons — have a positive (+) charge. Found in the nucleus of the atom.

The number of protons determines what element an atom is.

Example: Carbon has 6 protons

Atomic Number - the number of protons, each element has it own unique number of protons.

 Neutrons — a neutral charge (no charge) found in the nucleus. These determine whether an element is radioactive.

Protons and neutrons are about ovar atomic mass unit (annu). Electrons have a much smaller mass — it takes almost 2000 electrons to equal 1 annu.

Atomic Mass - the total number of protons and neutrons.

Mass Number - Atomic Mass that is rounded.

#### To find just how many neutrons an atom has:

# neutrons = atomic mass - atomic #

Isotopes - atoms of the same element that have different number of neutrons.

Carbon-12 Carbon-14
6 protons 6 protons
6 neutrons 8 neutrons

NOTICE - the number of protons DOES NOT change if it is to be carbon -- only the number of neutrons.

3. Electrons - have a negative (-) charge. Found moving around the atom is a space called orbital's (shells). Those electrons farthest away from the nucleus are called <u>valence electrons</u>, and are involved in the formation of chemical bonds.

In an uncharged atom the number of protons is equal to the number of electrons.

lon - when an atom gains or loses an electron (has a charge)

If an atom gains electrons = negative (-) charge

If an atom loses electrons - positive (+) charge

#### How to find:

# protons - atomic #

# neutrons = atomic mass - atomic #

# electrons = atomic # — charge (if any)

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**RC Schank** 

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- 1. Understanding the eBook Pearson Chemistry Atomic Structure Study Guide Answers
  - The Rise of Digital Reading Pearson Chemistry Atomic Structure Study Guide Answers
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Pearson Chemistry Atomic Structure Study Guide Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Pearson Chemistry Atomic Structure Study Guide Answers
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Pearson Chemistry Atomic Structure Study Guide Answers
  - Personalized Recommendations
  - Pearson Chemistry Atomic Structure Study Guide Answers User Reviews and Ratings
  - Pearson Chemistry Atomic Structure Study Guide Answers and Bestseller Lists
- 5. Accessing Pearson Chemistry Atomic Structure Study Guide Answers Free and Paid eBooks
  - Pearson Chemistry Atomic Structure Study Guide Answers Public Domain eBooks
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  - Pearson Chemistry Atomic Structure Study Guide Answers Budget-Friendly Options
- 6. Navigating Pearson Chemistry Atomic Structure Study Guide Answers eBook Formats
  - o ePub, PDF, MOBI, and More
  - Pearson Chemistry Atomic Structure Study Guide Answers Compatibility with Devices
  - Pearson Chemistry Atomic Structure Study Guide Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Pearson Chemistry Atomic Structure Study Guide Answers
  - Highlighting and Note-Taking Pearson Chemistry Atomic Structure Study Guide Answers
  - Interactive Elements Pearson Chemistry Atomic Structure Study Guide Answers
- 8. Staying Engaged with Pearson Chemistry Atomic Structure Study Guide Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Pearson Chemistry Atomic Structure Study Guide Answers
- 9. Balancing eBooks and Physical Books Pearson Chemistry Atomic Structure Study Guide Answers
  - Benefits of a Digital Library
  - o Creating a Diverse Reading Collection Pearson Chemistry Atomic Structure Study Guide Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Pearson Chemistry Atomic Structure Study Guide Answers
  - Setting Reading Goals Pearson Chemistry Atomic Structure Study Guide Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pearson Chemistry Atomic Structure Study Guide Answers
  - Fact-Checking eBook Content of Pearson Chemistry Atomic Structure Study Guide Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - o Integration of Multimedia Elements
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