

## Reviewing Content

42. The solvent is the substance in which the solute is dissolved.
43. Random collisions of the solvent molecules with the solute particles provide enough force to overcome gravity.
44. **solubility**: the amount of a substance that dissolves in a given quantity of solvent at specified conditions of temperature and pressure to produce a saturated solution.  
**saturated solution**: a solution containing the maximum amount of solute for a given amount of solvent at a constant temperature and pressure. **unsaturated solution**: a solution that contains less solute than a saturated solution at a given temperature and pressure. **miscible**: describes liquids that dissolve in each other. **immiscible**: describes liquids that are insoluble in each other.
45. Particles of solute crystallize.
46. No; if there were undissolved solute, the excess solute would come out of a supersaturated solution.
47.  $5.55 \times 10^2$  g  $\text{AgNO}_3$
48. Solubility increases with pressure.
49. a.  $1.6 \times 10^{-2}$  g/L  
b.  $4.7 \times 10^{-2}$  g/L
50. *Dilute* and *concentrated* are relative terms and are not quantitative. Molarity provides the exact number of moles of solute per liter of solution.
51. Molarity is the number of moles of solute dissolved in one liter of solution.  
a. 1.3M KCl  
b.  $3.3 \times 10^{-1}$ M  $\text{MgCl}_2$
52.  $2.00 \times 10^3$  mL
53. a.  $5.0 \times 10^{-3}$  mol NaCl, 29 g NaCl  
b. 1.0 mol  $\text{KNO}_3$ ,  $1.0 \times 10^{-2}$  g  $\text{KNO}_3$   
c.  $2.5 \times 10^3$  mol  $\text{CaCl}_2$ , 2.8 g  $\text{CaCl}_2$
54. a.  $2.3 \times 10^1$  g NaCl  
b. 2.0 g  $\text{MgCl}_2$
55. a. 16% (v/v) ethanol  
b. 63.6% (v/v) isopropyl alcohol
56. Colligative properties are properties of a solution that depend only on the number of solute particles; boiling-point elevation, freezing-point depression, and vapor-pressure lowering. Boiling points are elevated because shells of solvent form around solute particles, reducing the amount of solvent molecules that have sufficient energy to escape the solution; relative to the pure solvent, the amount of energy required to cause vaporization or boiling increases. Solutes disrupt the ordering of the solvent structure, so more kinetic energy must be withdrawn from a solution for it to solidify. This lowers the freezing point of the solution.
57. a. sea water  
b. 1.5M  $\text{KNO}_3$   
c. 0.100M  $\text{MgCl}_2$
58. The effective molality of the  $\text{Ca}(\text{NO}_3)_2$  solution is 3m. The effective molality of the  $\text{NaNO}_3$  solution is 2m.
59. When vapor pressure is lowered relative to pure solvent, more energy must be supplied to reach the boiling point; thus the boiling point is increased relative to pure solvent.
60. The salt lowers the freezing point of the ice-water cooling mixture.
61. 1M solution: 1 mol of solute in 1 L of solution; 1m solution: 1 mol of solute in 1000 g of solvent
62. Add 27.0 g  $\text{H}_2\text{O}$  to 32.0 g  $\text{CH}_3\text{OH}$ .
63. a.  $100.26^\circ\text{C}$   
b.  $101.54^\circ\text{C}$
64. a.  $-4.46^\circ\text{C}$   
b.  $-2.2^\circ\text{C}$
65. a.  $-1.1^\circ\text{C}$   
b.  $-0.74^\circ\text{C}$   
c.  $-1.5^\circ\text{C}$

## Understanding Concepts

66. a. The freezing-point depression is twice as great for solute B; solute B must provide twice as many particles in solution.

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**David C. Shelledy, Jay I. Peters**



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1. Understanding the eBook Pearson Chemistry Section 16 Review Answers
  - The Rise of Digital Reading Pearson Chemistry Section 16 Review Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Pearson Chemistry Section 16 Review Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Pearson Chemistry Section 16 Review Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Pearson Chemistry Section 16 Review Answers
  - Personalized Recommendations
  - Pearson Chemistry Section 16 Review Answers User Reviews and Ratings
  - Pearson Chemistry Section 16 Review Answers and Bestseller Lists
5. Accessing Pearson Chemistry Section 16 Review Answers Free and Paid eBooks
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  - Pearson Chemistry Section 16 Review Answers Enhanced eBook Features
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  - Interactive Elements Pearson Chemistry Section 16 Review Answers
8. Staying Engaged with Pearson Chemistry Section 16 Review Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Pearson Chemistry Section 16 Review Answers
9. Balancing eBooks and Physical Books Pearson Chemistry Section 16 Review Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Pearson Chemistry Section 16 Review Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Pearson Chemistry Section 16 Review Answers
  - Setting Reading Goals Pearson Chemistry Section 16 Review Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Pearson Chemistry Section 16 Review Answers
  - Fact-Checking eBook Content of Pearson Chemistry Section 16 Review Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

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