

Table 1.1 Relationship between Mole, Molar mass and Avogadro's Number			
Name of substance	Symbol/Formula	Molar Mass (g/mol)	Kind and number of particles in one mole
Sodium	Na	23.0	6.02×10^{23} Na atoms
Water	H ₂ O	18.0	6.02×10^{23} H ₂ O molecules 6.02×10^{23} O atoms $2(6.02 \times 10^{23})$ H atoms
Common Salt	NaCl	58.5	6.02×10^{23} NaCl Formula units 6.02×10^{23} Na ⁺ ions 6.02×10^{23} Cl ⁻ ions
Calcium Chloride	CaCl ₂	111	6.02×10^{23} CaCl ₂ Formula units 6.02×10^{23} Ca ²⁺ ions $2(6.02 \times 10^{23})$ Cl ⁻ ions
Ferric Chloride	FeCl ₃	162.35	6.02×10^{23} FeCl ₃ Formula units 6.02×10^{23} Fe ³⁺ ions $3(6.02 \times 10^{23})$ Cl ⁻ ions

Inter conversion of Mole and Mass

While working on stoichiometric problems, we often need to interconvert mole and mass of reactants and products. For this purpose molar mass is used as conversion factor. To convert given mass of substance into mole, we divide it by molar mass. On the other hand if moles are needed to be converted into mass, we simply multiply it with molar mass. Conclusively, the following formula may be used for the inter conversion of mole and mass.

$$\text{No. of moles} = \frac{\text{Given mass (g)}}{\text{Molar mass of substance}}$$

Example 1.1

Calculate the number of moles in 25.5g of sodium metal.

Data:

Given mass of sodium metal = 25.5 g

No. of moles of sodium metal = ?

Solution:

Since the molar mass of sodium is 23g/mol, we use this molar mass as conversion factor to determine the number of moles of sodium.

$$\text{Moles of Na} = \frac{\text{Given mass of Na}}{\text{Molar mass of Na}} = \frac{25.5\text{g}}{23\text{g/mol}} = 1.11 \text{ moles}$$

Pearson Chemistry Workbook Chapter 1stoichiometry

Liying Dong



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IIT Chemistry: Introductory Topics, **Iit Foundations - Chemistry Class 8** Pearson, 2009-09 IIT Foundation series is specifically for students preparing for IIT right from school days The series include books from class 8 to class 10th in physics chemistry mathematics **The Pearson Guide to Physical Chemistry for the IIT JEE** Singhal, **Teaching Science for Understanding** James Joseph Gallagher, 2007 Offers middle and high school science teachers practical advice on how they can teach their students key concepts while building their understanding of the subject through various levels of learning activities **Chemistry and Industry**, 1954 **Fundamentals of Water Treatment Unit Processes** David Hendricks, 2010-11-09 Carefully designed to balance coverage of theoretical and practical principles Fundamentals of Water Treatment Unit Processes delineates the principles that support practice using the unit processes approach as the organizing concept The author covers principles common to any kind of water treatment for example drinking water municipal wastewater industrial water treatment industrial waste water treatment and hazardous wastes Since technologies change but principles remain constant the book identifies strands of theory rather than discusses the latest technologies giving students a clear understanding of basic principles they can take forward in their studies Reviewing the historical development of the field and highlighting key concepts for each unit process each chapter follows a general format that consists of process description history theory practice problems references and a glossary This organizational style facilitates finding sections of immediate interest without having to page through an excessive amount of material Pedagogical Features End of chapter glossaries provide a ready reference and add terms pertinent to topic but beyond the scope of the chapter Sidebars sprinkled throughout the chapters present the lore and history of a topic enlarging students perspective Example problems emphasize tradeoffs and scenarios rather than single answers and involve spreadsheets Reference material includes several appendices and a quick reference spreadsheet Solutions manual includes spreadsheets for problems Supporting material is available for download Understanding how the field arrived at its present state of the art places the technology in a more logical context and gives students a strong foundation in basic principles This book does more than build technical proficiency it adds insight and understanding to the broader aspects of water treatment unit processes

Chemical Engineering Catalog, 1929 *Elements of Inorganic Structural Chemistry* Erwin Parthé, 1996 Physical Chemistry David S. Eisenberg, Donald M. Crothers, 1979 Books in Print Supplement, 1985 **Chemistry: The Central Science** Theodore L. Brown, H. Eugene LeMay Jr., Bruce E. Bursten, Catherine Murphy, Patrick Woodward, Steven Langford, Dalius Sagatys, Adrian George, 2013-10-04 If you think you know the Brown LeMay Bursten Chemistry text think again In response to market request we have created the third Australian edition of the US bestseller Chemistry The Central Science An extensive revision has taken this text to new heights Triple checked for scientific accuracy and consistency this edition is a more seamless and cohesive product yet retains the clarity innovative pedagogy functional problem solving and

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