

Chemical Reactions and Stoichiometry

1. Given the equation $3A + B \rightarrow C + D$, you react 1 mole of A with 3 moles of B. True or false: B is the limiting reactant because you have fewer moles of B than A.

Answer: False

$$1 \text{ mole A} \frac{1 \text{ mole B}}{3 \text{ mole A}} = 0.3 \text{ moles B needed. Have 3 moles B so A is limiting}$$

2. NaHCO_3 is the active ingredient in baking soda. How many grams of oxygen are in 0.35 g of NaHCO_3 ?

- a) 0.012 g
- b) 0.022 g
- c) 0.067 g
- d) 0.20 g

ANS: d)

$$0.35 \text{ g NaHCO}_3 \left(\frac{1 \text{ mole NaHCO}_3}{84.07 \text{ g NaHCO}_3} \right) \left(\frac{3 \text{ moles O}}{1 \text{ mole NaHCO}_3} \right) \left(\frac{16.00 \text{ g}}{1 \text{ mole O}} \right) = 0.199 = \boxed{0.20 \text{ g}}$$

3. Suppose the reaction $\text{Ca}_3(\text{PO}_4)_2 + 3\text{H}_2\text{SO}_4 \rightarrow 3\text{CaSO}_4 + 2\text{H}_3\text{PO}_4$ is carried out starting with 103 g of $\text{Ca}_3(\text{PO}_4)_2$ and 75.0 g of H_2SO_4 . How much phosphoric acid will be produced?

- a) 74.9 g
- b) 50.0 g
- c) 112 g
- d) 32.5 g

ANS: b)

$$\begin{aligned} 103 \text{ g Ca}_3(\text{PO}_4)_2 & \frac{1 \text{ mole}}{310.1875 \text{ g Ca}_3(\text{PO}_4)_2} = 0.332 \text{ moles Ca}_3(\text{PO}_4)_2 \\ 75.0 \text{ g H}_2\text{SO}_4 & \frac{1 \text{ mole H}_2\text{SO}_4}{98.082 \text{ g H}_2\text{SO}_4} = 0.765 \text{ moles H}_2\text{SO}_4 \text{ L.R.} \end{aligned}$$

$$\text{Ca}_3(\text{PO}_4)_2 \text{ needed: } 0.765 \text{ moles H}_2\text{SO}_4 \frac{1 \text{ mole Ca}_3(\text{PO}_4)_2}{3 \text{ moles H}_2\text{SO}_4} = 0.255 \text{ moles Have excess}$$

$$0.765 \text{ moles H}_2\text{SO}_4 \frac{2 \text{ moles H}_3\text{PO}_4}{3 \text{ moles H}_2\text{SO}_4} \frac{97.994 \text{ g H}_3\text{PO}_4}{1 \text{ mole}} = 49.97 \text{ g} = \boxed{50.0 \text{ g}}$$

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G Orfield



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