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Ions in Aqueous Solutions and Colligative Properties

Section Quiz: Colligative Properties of Solutions

Write your answer provided with the letter of the best or phrase that best completes each statement or best answers each question.

- _____ 1. Which is/are not colligative properties of solutions?
- It is directly proportional to the number of particles in solution.
 - It is directly proportional to the number of particles in solution.
 - It depends on the properties of an individual solute molecule.
 - Answer all three above.
- _____ 2. Compounds with a total of 4 Cl^- , 2 Ca^{2+} , and 10 Na^+ ions in a solution give:
- the same freezing-point depression.
 - about twice the freezing-point depression.
 - the same freezing-point elevation.
 - about one twice the freezing-point elevation.
- _____ 3. Compound with a total of four particles, a total of 10 particles, and a total of 2 particles in a solution give:
- the same freezing-point depression.
 - about half of the freezing-point depression.
 - a freezing-point elevation about two-thirds of 1/3.
 - a freezing-point elevation about 1/3 of one of 1/3.
- _____ 4. When a nonelectrolyte solute dissolved in water, the liquid portion of the solution:
- increases.
 - decreases.
 - does not change.
 - changes depending on the solute used.
- _____ 5. Which of the following compounds would be most effective in lowering the boiling point of 100 mL water?
- $\text{C}_2\text{H}_6\text{O}$
 - H_2SO_4
 - H_2O
 - H_2SO_4

Rosengarten Colligative Properties Homework Answers

G Thomas



Rosengarten Colligative Properties Homework Answers:

Decoding **Rosengarten Colligative Properties Homework Answers**: Revealing the Captivating Potential of Verbal Expression

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