

AP Chemistry – Chapters 12 - 14 Review Questions

Vocab/Concepts to Review:

- What is a boiling point? What is the “normal boiling point”?
 - Which state changes are endothermic/exothermic?
 - What is a supercritical fluid (SCF) and what are its properties?
 - What are the different packing structures of solid? (simple cubic, fcc, bcc)
 - Know how to read a phase diagram!
 - What are amorphous solids? Give some examples.
 - What are colligative properties?
 - What determines the solubility of gases in a solvent?
1. Which of the following molecular solids will exhibit dipole-dipole intermolecular forces: NH_3 , BF_3 , I_2 , and H_2S ?
 2. How can you compare boiling points among molecules? **stronger intermolecular forces causes higher boiling points**
 3. Which is the dominant intermolecular force present in acetic acid, $\text{CH}_3\text{CO}_2\text{H}(\text{l})$? **H-bonds**
 4. Which intermolecular forces are present in $\text{CH}_3\text{F}(\text{s})$? **LDF and dipole-dipole**
 5. What properties of water can be attributed to hydrogen bonding? **increased melting point, increased heat of vaporization, decreased vapor pressure, increased surface tension**
 6. At its boiling point of -10.0°C , 20.7 kJ of heat is required to vaporize 53.2 g of sulfur dioxide. What is the enthalpy of vaporization of SO_2 ? **24.9 kJ/mol**
 7. Which process requires the greatest exothermic change in enthalpy for water? **condensation**
 8. At 25°C , water has an equilibrium vapor pressure of 23.8 mm Hg. If 2.500 g of water is sealed in an evacuated 10.0-L flask and heated to 25°C , what mass of water will evaporate? Assume any liquid remaining in the flask has a negligible volume. ($R = 0.08206 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K}$) **0.231 grams**
 9. What is the definition of molality? **mol solute/kg solvent**
 10. What is the molality of Na^+ ions in a 6.29% by mass Na_2CO_3 solution? The molar mass of sodium carbonate is 105.99 g/mol. **1.27 m**
 11. What is the mole fraction of KBr present in a solution that is 11.04% by mass aqueous KBr? The molar mass of KBr is 119.0 g/mol. **0.01845**
 12. Concentrated NaOH is 19.3 M and has a density of 1.53 g/mL. What is the percent NaOH by mass of concentrated NaOH? **50.5%**
 13. The Henry's law constant for N_2 in water at 25°C is $8.4 \times 10^{-7} \text{ M}/\text{mm Hg}$. What is the equilibrium concentration of N_2 in water when the partial pressure of N_2 is 653 mm Hg? **$5.5 \times 10^{-4} \text{ M}$**
 14. Equal mass of water and ethylene glycol are mixed. What is the vapor pressure of this mixture at 100°C ? The molar masses of water and ethylene glycol are 18.02 g/mol and 62.07 g/mol, respectively. Assume ideal behavior for the solution. **589 mm Hg (this is a tricky one!)** You are going to be using Raoult's Law.
 15. What is the freezing point of a solution containing 5.663 grams naphthalene (molar mass = 128 g/mol) dissolved in 32.0 grams of paradichlorobenzene? The freezing point of pure paradichlorobenzene is 53.0°C and the freezing point depression constant, K_{fp} , is $-7.10^\circ\text{C}/\text{m}$. **-43.2°C**
 16. What is the molar mass of a compound if 5.12 grams is dissolved in 25.0 grams of chloroform solvent to form a solution that has a boiling point of 67.54°C ? The boiling point of pure chloroform is 61.70°C and the boiling point elevation constant, K_{bp} , is $3.63^\circ\text{C}/\text{m}$. **127 g/mol**
 17. The osmotic pressure of blood is 7.65 atm at 37°C . What mass of glucose ($\text{C}_6\text{H}_{12}\text{O}_6$, molar mass = 180.2 g/mol) is needed to prepare 2.00 L of solution for intravenous injection? The osmotic pressure of the glucose solution must equal the osmotic pressure of blood. **108 grams**

Pearson Chemistry Chapters 12 Review Answers

Kamthane Ashok



Pearson Chemistry Chapters 12 Review Answers:

Introduction to Geochemistry Kula C. Misra, 2012-03-28 INTRODUCTION TO Geochemistry This book is intended to serve as a text for an introductory course in geochemistry for undergraduate graduate students with at least an elementary level background in earth sciences chemistry and mathematics The text containing 83 tables and 181 figures covers a wide variety of topics ranging from atomic structure to chemical and isotopic equilibria to modern biogeochemical cycles which are divided into four interrelated parts Crystal Chemistry Chemical Reactions and biochemical reactions involving bacteria Isotope Geochemistry radiogenic and stable isotopes and The Earth Supersystem which includes discussions pertinent to the evolution of the solid Earth the atmosphere and the hydrosphere In keeping with the modern trend in the field of geochemistry the book emphasizes computational techniques by developing appropriate mathematical relations solving a variety of problems to illustrate application of the mathematical relations and leaving a set of questions at the end of each chapter to be solved by students However so as not to interrupt the flow of the text involved chemical concepts and mathematical derivations are separated in the form of boxes Supplementary materials are packaged into ten appendixes that include a standard state 298 15 K 1 bar thermodynamic data table and a listing of answers to selected chapter end questions

Sif: Chemistry 5na Tb J. G. R. Briggs, 2009 *Chemistry: The Central Science* Theodore L. Brown, H. Eugene LeMay Jr., Bruce E. Bursten, Catherine Murphy, Patrick Woodward, Steven Langford, Dalius Sagatys, Adrian George, 2013-10-04 If you think you know the Brown LeMay Bursten Chemistry text think again In response to market request we have created the third Australian edition of the US bestseller Chemistry The Central Science An extensive revision has taken this text to new heights Triple checked for scientific accuracy and consistency this edition is a more seamless and cohesive product yet retains the clarity innovative pedagogy functional problem solving and visuals of the previous version All artwork and images are now consistent in quality across the entire text And with a more traditional and logical organisation of the Organic Chemistry content this comprehensive text is the source of all the information and practice problems students are likely to need for conceptual understanding development of problem solving skills reference and test preparation *Descriptive Inorganic Chemistry* James E. House, Kathleen A. House, 2010-09-22 Descriptive Inorganic Chemistry Second Edition covers the synthesis reactions and properties of elements and inorganic compounds for courses in descriptive inorganic chemistry This updated version includes expanded coverage of chemical bonding and enhanced treatment of Buckminster Fullerenes and incorporates new industrial applications matched to key topics in the text It is suitable for the one semester ACS recommended course or as a supplement in general chemistry courses Ideal for majors and non majors the book incorporates rich graphs and diagrams to enhance the content and maximize learning Includes expanded coverage of chemical bonding and enhanced treatment of Buckminster Fullerenes Incorporates new industrial applications matched to key topics in the text

The School Science Review, 2002 **Inorganic Chemistry** Catherine E. Housecroft, A. G. Sharpe, 2008 Designed as a

student text Inorganic Chemistry focuses on teaching the underlying principles of inorganic chemistry in a modern and relevant way

Phlebotomy Essentials Ruth E. McCall, 2023-08-11 Phlebotomy Essentials Eighth Edition provides accurate up to date and practical information and instruction in phlebotomy procedures and techniques along with a comprehensive background in phlebotomy theory and principles It is appropriate for use as an instructional text or as a reference for those who wish to update skills or study for national certification This updated edition details how today s phlebotomists work in an approach optimized for how today s students learn Combined with an optional Workbook Exam Review book and updated digital courseware the latest edition of Phlebotomy Essentials represents a cornerstone of preparation for a successful career in phlebotomy Complete program integration of textbook workbook and exam review products working together as a complete learning system for optimal student outcomes Phlebotomy Essentials Eighth Edition with Navigate Premier Access Student Workbook for Phlebotomy Essentials Eighth Edition Phlebotomy Exam Review Eighth Edition with Navigate Premier Access Phlebotomy Essentials Updated to current CLSI guidelines TestPrep alerts throughout the textbook share current statistics on which questions students typically get wrong and why Map of NAACLS Standards where each competency is covered Certification Test Plan Alignment Map including American Medical Technologists AMT Registered Phlebotomy Technician American Society for Clinical Pathology ASCP Phlebotomy Technician National Healthcareer Association NHA Certified Phlebotomy Technician National Center for Competency Testing NCCT National Certified Phlebotomy Technician 2024 516 pages

Respiratory Care: Patient Assessment and Care Plan Development David C. Shelledy, Jay I. Peters, 2021-02-08 Respiratory Care Patient Assessment and Care Plan Development Second Edition describes the purpose of patient assessment and then guides the reader through the process of reviewing existing data in the medical record

Introductory Chemistry Nivaldo J. Tro, 2006

The Cambridge Review, 1895 Vols 1 26 include a supplement The University pulpit vols 1 26 no 1 661 which has separate pagination but is indexed in the main vol

C Programming: Test Your Skills Kamthane Ashok, 2010-09 C Programming Test Your Skills is specifically designed to be used as the supplementary resource for learning C Programming It is ideal for self practice or test preparation and hones one s problem solving abilities through varieties of exercises

Chemistry 'O' Level Rex M. Heyworth, 2007

Chemistry Catherine E. Housecroft, Edwin C. Constable, 2010 This text integrates the three major branches of chemistry with the aim of enabling students to tackle more easily the problems within the subject and to apply chemistry to real life situations

Fundamentals of Biochemistry Donald Voet, Judith G. Voet, Charlotte W. Pratt, 2016-02-29 Voet Voet and Pratt s Fundamentals of Biochemistry 5th Edition addresses the enormous advances in biochemistry particularly in the areas of structural biology and Bioinformatics by providing a solid biochemical foundation that is rooted in chemistry to prepare students for the scientific challenges of the future While continuing in its tradition of presenting complete and balanced coverage that is clearly written and relevant to human health and disease Fundamentals of Biochemistry 5e includes new

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India Rubber World and Electrical Trades Review John Robertson Dunlap, Henry Clemens Pearson, 1950

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Series ICSE Chemistry 7 Lorraine Mirza, Krishnakali Gupta, **150 technical questions and answers for job interview Offshore Drilling Rigs** Petrogav International Oil & Gas Training Center, 2020-06-28 The job interview is probably the most important step you will take in your job search journey Because it s always important to be prepared to respond effectively to the questions that employers typically ask at a job interview Petrogav International has prepared this eBooks that will help you to get a job in oil and gas industry Since these questions are so common hiring managers will expect you to be able to answer them smoothly and without hesitation This eBook contains 150 questions and answers for job interview and as a BONUS 230 links to video movies This course covers aspects like HSE Process Mechanical Electrical and Instrumentation Control that will enable you to apply for any position in the Oil and Gas Industry

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