In the list below, circle the properties of acids and box the properties of bases.

Sour taste Feel slippery Contains hydroxide ions

Turns litmus blue Reacts with metals to form Ho Turns litmus red

Bitter taste Contains hydrogen ions Corrosive

Complete the chart below:

| Acid-Base Definitions |                               |                               |
|-----------------------|-------------------------------|-------------------------------|
| Type                  | Acid Definition               | Base Definition               |
| Arrhenius             | Donates an H' ion in solution | Donates an OH ion in solution |
| Bronsted-Lowry        | Donates an H* ion             | Accepts an H* ion             |
| Lewis                 | Accepts an electron pair      | Donates an electron pair      |

- What term(s) describe a substance that can act as an acid or a base? Amphoteric or amphiprotic
- 4. Neutralization reactions always produce a SALT and WATER
- If the [H<sup>\*</sup>] is greater than the [OH<sup>\*</sup>], the solution is acidic basic neutral
- 6. If the [OH] is greater than the [H], the solution is acidic basic neutral
- 7. If the [H'] is equal to the [OH'], the solution is acidic basic neutral
- 8. If the pH is 7, the solution is acidic basic neutral
- If the pH is greater than 7, the solution is acidic basic neutral
- 10. If the pH is less than 7, the solution is acidic basic neutral
- 11. How many moles of lithium hydroxide are required to neutralize 19.2 grams of chloric acid?

(19.2 g HClO<sub>2</sub>) x (1 mol HClO<sub>2</sub>/84 g HClO<sub>2</sub>) x (1 mol LiOH/1 mol HClO<sub>2</sub>) = 0.228 mol LiOH

12. How many grams of calcium hydroxide are required to neutralize 132.4 grams of nitric acid?

 $Ca(OH)_2 + 2 HNO_3 \rightarrow Ca(NO_3)_2 + 2 H_2O$ 

(132.4 g HNO<sub>3</sub>) x (1 mol HNO<sub>3</sub>/63 g HNO<sub>3</sub>) x (1 mol Ca(OH)<sub>2</sub>/2 mol HNO<sub>3</sub>) x (74 g Ca(OH)<sub>2</sub>/1 mol Ca(OH)<sub>2</sub>) = <u>77.76 g Ca(OH)<sub>3</sub></u>

# **Pearson Chemistry Work Answers Chapter Acids Bases**

James E. House, Kathleen A. House

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## **Table of Contents Pearson Chemistry Work Answers Chapter Acids Bases**

- 1. Understanding the eBook Pearson Chemistry Work Answers Chapter Acids Bases
  - The Rise of Digital Reading Pearson Chemistry Work Answers Chapter Acids Bases
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Pearson Chemistry Work Answers Chapter Acids Bases
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Pearson Chemistry Work Answers Chapter Acids Bases
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Pearson Chemistry Work Answers Chapter Acids Bases

- Personalized Recommendations
- Pearson Chemistry Work Answers Chapter Acids Bases User Reviews and Ratings
- Pearson Chemistry Work Answers Chapter Acids Bases and Bestseller Lists
- 5. Accessing Pearson Chemistry Work Answers Chapter Acids Bases Free and Paid eBooks
  - Pearson Chemistry Work Answers Chapter Acids Bases Public Domain eBooks
  - Pearson Chemistry Work Answers Chapter Acids Bases eBook Subscription Services
  - Pearson Chemistry Work Answers Chapter Acids Bases Budget-Friendly Options
- 6. Navigating Pearson Chemistry Work Answers Chapter Acids Bases eBook Formats
  - o ePub, PDF, MOBI, and More
  - Pearson Chemistry Work Answers Chapter Acids Bases Compatibility with Devices
  - Pearson Chemistry Work Answers Chapter Acids Bases Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Pearson Chemistry Work Answers Chapter Acids Bases
  - Highlighting and Note-Taking Pearson Chemistry Work Answers Chapter Acids Bases
  - Interactive Elements Pearson Chemistry Work Answers Chapter Acids Bases
- 8. Staying Engaged with Pearson Chemistry Work Answers Chapter Acids Bases
  - o Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Pearson Chemistry Work Answers Chapter Acids Bases
- 9. Balancing eBooks and Physical Books Pearson Chemistry Work Answers Chapter Acids Bases
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Pearson Chemistry Work Answers Chapter Acids Bases
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Pearson Chemistry Work Answers Chapter Acids Bases
  - Setting Reading Goals Pearson Chemistry Work Answers Chapter Acids Bases
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pearson Chemistry Work Answers Chapter Acids Bases

- Fact-Checking eBook Content of Pearson Chemistry Work Answers Chapter Acids Bases
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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