

1. In the list below, circle the properties of acids and box the properties of bases.

Sour taste	Feel slippery	Contains hydroxide ions
Turns litmus blue	Reacts with metals to form $H_2$	Turns litmus red
Bitter taste	Contains hydrogen ions	Corrosive

2. Complete the chart below:

Acid-Base Definitions		
Type	Acid Definition	Base Definition
Arrhenius	Donates an $H^+$ ion in solution	Donates an $OH^-$ ion in solution
Bronsted-Lowry	Donates an $H^+$ ion	Accepts an $H^+$ ion
Lewis	Accepts an electron pair	Donates an electron pair

3. What term(s) describe a substance that can act as an acid or a base? **Amphoteric or amphiprotic**
4. Neutralization reactions always produce a **SALT** and **WATER**
5. If the  $[H^+]$  is greater than the  $[OH^-]$ , the solution is acidic basic neutral
6. If the  $[OH^-]$  is greater than the  $[H^+]$ , the solution is acidic basic neutral
7. If the  $[H^+]$  is equal to the  $[OH^-]$ , the solution is acidic basic neutral
8. If the pH is 7, the solution is acidic basic neutral
9. If the pH is greater than 7, the solution is acidic basic neutral
10. If the pH is less than 7, the solution is acidic basic neutral
11. How many moles of lithium hydroxide are required to neutralize 19.2 grams of chloric acid?



$$(19.2 \text{ g } HClO_3) \times (1 \text{ mol } HClO_3 / 84 \text{ g } HClO_3) \times (1 \text{ mol } LiOH / 1 \text{ mol } HClO_3) = \underline{0.228 \text{ mol } LiOH}$$

12. How many grams of calcium hydroxide are required to neutralize 132.4 grams of nitric acid?



$$(132.4 \text{ g } HNO_3) \times (1 \text{ mol } HNO_3 / 63 \text{ g } HNO_3) \times (1 \text{ mol } Ca(OH)_2 / 2 \text{ mol } HNO_3) \times (74 \text{ g } Ca(OH)_2 / 1 \text{ mol } Ca(OH)_2) = \underline{77.76 \text{ g } Ca(OH)_2}$$

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*Chemical Demonstrations* Bassam Z. Shakhashiri, 1983. The demonstrations capture interest, teach, inform, fascinate, amaze and perhaps most importantly involve students in chemistry. Nowhere else will you find books that answer: How come it happens? Is it safe? What do I do with all the stuff when the demo is over? Shakhashiri and his collaborators offer 282 chemical demonstrations arranged in 11 chapters. Each demonstration includes seven sections: a brief summary, a materials list, a step by step account of procedures to be used, an explanation of the hazards involved, information on how to store or dispose of the chemicals used, a discussion of the phenomena displayed and principles illustrated by the demonstration, and a list of references. You will find safety emphasized throughout the book in each demonstration.

*Handbook of Hydrothermal Technology* K. Byrappa, Masahiro Yoshimura, 2008-12-10. Quartz, zeolites, gemstones, perovskite type oxides, ferrite, carbon allotropes, complex coordinated compounds and many more all products now being produced using hydrothermal technology. *Handbook of Hydrothermal Technology* brings together the latest techniques in this rapidly advancing field in one exceptionally useful, long needed volume. The handbook provides a single source for understanding how aqueous solvents or mineralizers work under temperature and pressure to dissolve and recrystallize normally insoluble materials and decompose or recycle any waste material. The result, as the authors show in the book, is technologically the most efficient method in crystal growth, materials processing and waste treatment. The book gives scientists and technologists an overview of the entire subject including: Evolution of the technology from geology to widespread industrial use; Descriptions of equipment used in the process and how it works; Problems involved with the growth of crystals; processing of technological materials; environmental and safety issues; Analysis of the direction of today's technology. In addition, readers get a close look at the hydrothermal synthesis of zeolites, fluorides, sulfides, tungstates and molybdates, as well as native elements and simple oxides. Delving into the commercial production of various types, the authors clarify the effects of temperature, pressure, solvents and various other chemical components on the hydrothermal processes.

*Basic Analytical Chemistry* L. Pataki, E. Zapp, 2013-10-22. Pergamon Series in Analytical Chemistry Volume 2. *Basic Analytical Chemistry* brings together numerous studies of the vast expansion in the use of classical and instrumental methods of analysis. This book is composed of six chapters. After providing a theoretical background of analytical chemistry, this book goes on dealing with the fundamental principles of chemical equilibria in solution. The subsequent chapters consider the advances in qualitative and quantitative chemical analyses. These chapters present a unified view of these analyses based on the Bronsted-Lowry theory and the donor

acceptor principle These topics are followed by discussions on instrumental analysis using various methods including electrochemical optical spectroscopic and thermal methods as well as radioactive isotopes The final chapters examine the separation methods and the essential features of organic chemical analysis that are different from methods for inorganic compounds This book is of value to analytical chemists and researchers

**Hard and Soft Acids and Bases** Ralph G. Pearson, 1973 *Longman Science Chemistry 10* Kohli Nitin, 2008-09

**Ionic Surfactants and Aqueous Solutions** Juan H. Vera, Grazyna Wilczek-Vera, 2018-07-09

**Ionic Surfactants and Aqueous Solutions** Biomolecules Metals and Nanoparticles covers a wide range of subjects related to aqueous systems from reverse micelles as ion exchangers to the study of micellar phase transfer catalysis for nucleophilic substitution reactions The diverse background expertise and professional interests of the contributors to this book give it a unique richness of approach in topics of relevance for biotechnology and environmental studies Over sixty publications presenting research results are combined and expanded in this book by some of the original researchers At a mature age and at the summit of successful professional careers they have taken a second look to the state of the art in the fields that they had pioneered Eva Rodil and Ana Soto who had their research formation in the group of Professor Alberto Arce at Universidad de Santiago de Compostela Spain are presently professors at that university Maen Husein is a professor at University of Calgary Canada Remy Dumortier Mohammad Khoshkbarchi Hamid Rabie and Younok Dumortier Shin are presently active leaders in the industrial world in Canada and the USA The editors are retired academics from McGill University Montreal Canada and coauthors of the book *Classical Thermodynamics of Fluid Systems*

**Soft Matter at Aqueous Interfaces** Peter Lang, Yi Liu, 2015-12-24 This book covers the science of interfaces between an aqueous phase and a solid another liquid or a gaseous phase starting from the basic physical chemistry all the way to state of the art research developments Both experimental and theoretical methods are treated thanks to the contributions of a distinguished list of authors who are all active researchers in their respective fields The properties of these interfaces are crucial for a wide variety of processes products and biological systems and functions such as the formulation of personal care and food products paints and coatings microfluidic and lab on a chip applications cell membranes and lung surfactants Accordingly research and expertise on the subject are spread over a broad range of academic disciplines and industrial laboratories This book brings together knowledge from these different places with the aim of fostering education collaborations and research progress

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*Transactions of the Pharmaceutical Meetings* ,1878

**Participation Act for the International Atomic Energy**

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Cooperation,United States. Congress. Joint Committee on Atomic Energy. Subcommittee on Research and Development,1957

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Metals Speciation Separation and Recovery James W. Patterson,1987-07-05 This

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