#### CHEM 105-Fundamentals of Chemistry

#### Problem Set 6-Thermodynamics

- Predict the sign (+ or -) of ΔS for the following processes:
- a.  $Br_2(g) \rightarrow 2Br(g)$
- H<sub>2</sub>O (I) → H<sub>2</sub>O (g)
- c. Sugar (s) → Sugar (aq)
- d. 2SO<sub>2</sub>(g) + O<sub>2</sub>(g) → 2SO<sub>3</sub>(g)
- e. Cooling of water from 70 °C to 14 °C
- Calculate ΔS when 1 g of H<sub>2</sub>O (I) is reversibly evaporated at 100 °C. (ΔH<sub>vap</sub>=540 cal/g)
- 3. Some assymetrical compounds like CO may not have third law of thermodynamics (The entropy of a pure perfect crystal at 0 K is zero). Can you explain why?
- Boiling point of propane is -22 °C at 1 atm. Calculate Δ H<sub>sup</sub> for propane using Trouton's rule.
- "Dissociation of a diatomic gas is favored by high temperature". Justify this statement by using thermodynamics. Use your knowledge of bond energy and entropy.
- 6. A traveller goes from California to Colorado. Which of the followings are state functions?
- a. distance traveled
- b. latitude change
- c. altitude change
- gasoline consumed
- e. elapsed time
- f. longitude change
- 7. Given the following information:  $\Delta H^{o}_{f}(H_{2}O)(g)=-242 \text{ kJ/mol}$  $\Delta H^{o}_{f}(CH_{4})(g)=-74.9 \text{ kJ/mol}$

the heat of combustion of CH<sub>4</sub> (g) to CO<sub>2</sub> (g) and H<sub>2</sub>O (g) is -803 kJ/mol.

- a. Calculate ΔH<sup>o</sup><sub>rss</sub> for the reaction: C (s) + 2H<sub>2</sub>O (g) → CO<sub>2</sub> (g) + 2H<sub>2</sub> (g)
- b. Calculate Δ U° for the same reaction
- Use thermodynamic data at 298 K to decide in which direction the below reaction is spontaneous when the partial pressures of SO<sub>2</sub> (g), O<sub>2</sub> (g) and SO<sub>3</sub> (g) are 1x10<sup>-4</sup>, 0.20 and 0.10 atm, respectively.

 $2SO_2(g) + O_2(g) = 2SO_3(g)$   $\Delta G^{\circ}_1(SO_2)(g) = -300.2 \text{ kJ/mol}$  $\Delta G^{\circ}_1(SO_3)(g) = -371.1 \text{ kJ/mol}$ 

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**David Baud** 

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