

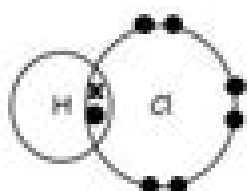
# C3 - STRUCTURE AND BONDING 3.5 Covalent bond

## Aiming for Grade 4

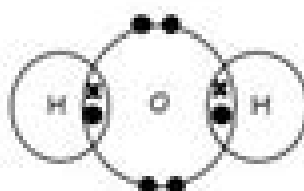
Ionic bonds form between **metals** and **non-metals**. Covalent bonds form between **non-metals**. Covalent bonds form when atoms **share** pairs of **electrons**. These bonds are very **strong**.

**Share, Metals, strong, non-metals, electrons**  
You can use the words more than once.

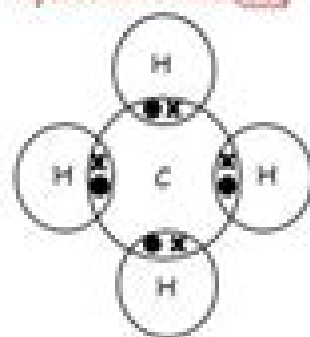
Name these small molecules and state their formulae.



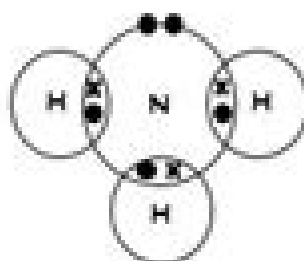
Hydrochloric Acid ( $\text{HCl}$ )



Water ( $\text{H}_2\text{O}$ )



Methane ( $\text{CH}_4$ )



Ammonia ( $\text{NH}_3$ )

Assessment

## Aiming for Grade 6

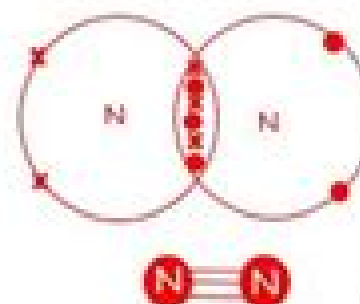
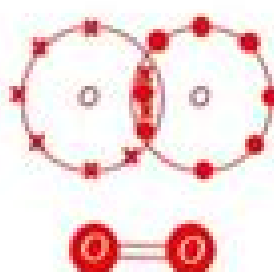
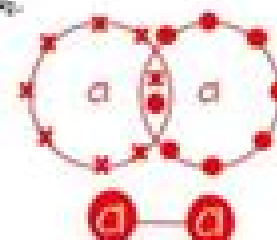
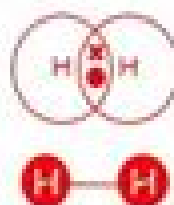
Describe how Hydrogen and Chlorine form a covalent bond to form Hydrochloric acid.

To gain full outer shells Hydrogen needs to gain an electron and Chlorine needs to lose one. Hydrogen shares its outer shell electron with Chlorine and Chlorine shares one of its electrons.

What is a double covalent bond?

Two shared pairs of electrons

Draw dot and cross diagrams and ball and stick diagrams for  $\text{H}_2$ ,  $\text{Cl}_2$ ,  $\text{O}_2$ , and  $\text{N}_2$ .



Assessment

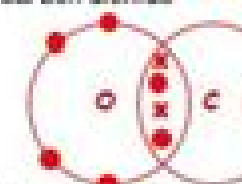
## Aiming

Draw ball and stick or the following covalent

Three hydrogen atoms  
one phosphorus (phos)



Carbon dioxide



Ethane (2 carbon with hydrogen)



How do double and triple properties?

A double or triple bond the bond making the me

Assessment  
Grade 6

1. I can draw dot and

# The Science Lab C3 C3 Mark Scheme

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Point Group Symmetry Applications Philip H. Butler, 2012-12-06 The mathematical apparatus of group theory is a means of exploring and exploiting physical and algebraic structure in physical and chemical problems The existence of structure in the physical processes leads to structure in the solutions For group theory to be useful this structure need not be an exact symmetry although as examples of exact symmetries we have that the identity of electrons leads to permutation symmetries in many electron wave functions the spatial structure of crystals leads to the Bloch theory of crystal eigenfunctions and the rotational invariance of the hydrogenic Hamiltonian leads to its factorization into angular and radial parts In the 1930 s Wigner extended what is known to mathematicians as the theory of group representations and the theory of group algebras to study the coupling coefficients of angular momentum relating various properties of the coefficients to the properties of the abstract group of rotations in 3 space In 1949 Racah in a paper on rare earth spectra showed that similar coefficients occur in other situations Immediately a number of studies of the coefficients were begun notably by Jahn with his applications in nuclear physics In the years since then a large number of physicists and chemists have added to the development of a general theory of the coefficients or have produced specialized tables for a specific application Applications now range from high energy physics to biology     **The Summary of Engineering Research** University of Illinois at Urbana-Champaign. Office of Engineering Publications, 1985

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