The properties of gases

1A The perfect gas

Answers to discussion questions

IA.2 The partial pressure of a gm in a minture of gmes is the pressure the gm would exect if it occupied alone the same contriber as the minture at the same temperature. Delitor's lary is a limiting law because it bolds exactly only moder conditions when the gases have no effect upon each other. This can only be true in the limit of zero pressure when the molecules of the gm are very far apart. Hence, Dalitor's law bolds exactly only for a minture of perfect gases, for real gases, the law is only an approximation.

Solutions to exercises

1A.1(b) The perfect gas law [1A.5] is pV = nRT, implying that the pressure would be

$$p = \frac{m\Omega T}{U}$$

All quantities on the right am given to us except e, which can be computed from the given mass of Ar.

$$w = \frac{25 \text{ g}}{39.95 \text{ g mol}^{-1}} = 0.626 \text{ mol}$$

so
$$p = \frac{(0.625 \text{ mol}) \times (8.31 \times 10^{-2} \text{dm}^2 \text{ ber K}^{-1} \text{ mol}^{-1}) \times (30 + 273) \text{ K}}{1.5 \text{ dm}^2} = \frac{10.5 \text{ ber}}{1.5 \text{ dm}^2}$$

So not the sample would not exert a pressure of 2.0 bar.

1A.2(b) Boyle's law [1A.4a] applies.

pV = constant so $p_tV_t = p_tV_t$ Solve for the initial pressure:

(0)
$$p_i = \frac{p_i V_c}{V_i} = \frac{(1.97 \text{ har}) \times (2.14 \text{ dem}^2)}{(2.14 + 1.80) \text{ dem}^3} = 1.07 \text{ har}$$

(II) The original persaure in Torr is

$$p_i = (1.07 \text{ barr}) \times \left(\frac{1 \text{ sten}}{1.013 \text{ barr}}\right) \times \left(\frac{760 \text{ Toor}}{1 \text{ sten}}\right) = 803 \text{ Toor}$$

1A.3(b) The relation between pressure and temperature at constant volume can be dedied from the perfect gas law, pV = nRT [1A.5]

as
$$p \propto T$$
 and $\frac{P_1}{T_1} = \frac{P_2}{T_1}$

The final pressure, then, could to be

$$P_c = \frac{P_c T_c}{T_c} = \frac{(125 \text{ kPa}) \times (11 + 273) \text{K}}{(23 + 273) \text{K}} = \boxed{120 \text{ kPa}}$$

1A.4(b) According to the perfect gas law [1.8], one can compute the amount of gas from pressure, temperature, and volume.
aV = uST

$$u = \frac{pV}{RT} = \frac{(1.00 \text{ atm.}) \times (1.013 \times 10^{9} \text{ Pa atm.}^{-1}) \times (4.00 \times 10^{9} \text{ cm}^{3})}{(8.3145 \text{ J.K.}^{-1} \text{mol}^{-1}) \times (20 + 273) \text{K.}} = 1.66 \times 10^{9} \text{ mol}$$

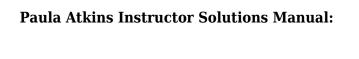
Once this is done, the mass of the gas can be computed from the amount and the molar mass:

$$m = (1.66 \times 10^{5} \text{mal}) \times (16.04 \text{ g mal}^{-1}) = 2.67 \times 10^{5} \text{ g} = 2.67 \times 10^{3} \text{ kg}$$

1A.5(b) The total pressure is the external pressure plus the hydrostatic pressure [1A.1], making the total pressure.

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