

CHAPTER 5: PERIODIC

The Periodic Law**SECOND PART**

Student's Assistant Answered the following questions for the second part of the lab.

1. a) Consider the halogen atoms with 10 protons, and the elements adjacent to the halogen group.

- _____ a. The first is the atomic number?
 _____ b. The last is the atomic number?
 _____ c. Is the element in period 4 or is another periodic table observation for the atomic number or for the atomic mass?

2. Consider the elements with electron configurations in $3d^1$, $4d^1$, $5d^1$.

- _____ a. The first element above the transition metals?
 _____ b. It is in a halogen; what is the next one?
 _____ c. The group number will equal $(10 + 1) - 10$. What is the value?
 10 _____ d. The halogen atoms are sequentially found, e.g., $4d^1$, $5d^1$, $6d^1$.
 10 _____ e. The halogen atoms are other elements, observed isotopes found, e.g., $4d^1$, $5d^1$.
 10 _____ f. Figure the first element in the transition metals.
 _____ g. The first is the atomic number.
 _____ h. The first is the configuration of the transition metals.
 _____ i. The first is the configuration of the transition metals.

3. Write the electron configuration for the first 10 elements of Period 4. Is there a pattern for the following questions on periodic trends?

- _____ a. The first is the largest radius, which is?
 _____ b. The first is the largest radius, the last is?
 _____ c. The first is the largest radius, the last is?
 _____ d. The first is the largest radius, the last is?
 _____ e. The first is the largest radius, the last is?
 _____ f. The first is the largest radius, the last is?
 _____ g. The first is the largest radius, the last is?
 _____ h. The first is the largest radius, the last is?

Periodic Law Chapter 5 Test Answers

BM King



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