

CHEMISTRY Objective Final Exam Review 10-11

There are 50 questions on the test. The chapters are listed below. Each of the letters represents a concept/idea that you should know or look up.

1. CH 6: Periodic trends, where do they increase & decrease
 - a. Atomic radius & Ionic radius
 - b. Electronegativity
 - c. Electron affinity
 - d. Ionization energy
 - e. Reactivity of metals and of non-metals
2. CH 7: Look up properties of elements per group and the names of each group.
 - a. Group 1 very reactive in water; group 7 most reactive non-metals etc.
 - b. Fluorine most reactive in periodic table
 - c. Group 1 will lose e- to become +1
 - d. Elements in same group have same chem properties.
3. CH 8: Ionic compounds and bonds
 - a. Why do atoms bond? How many electrons make them stable?
 - b. How to name ionic compounds, go from name to formula
 - c. What elements make up ionic compounds?
 - d. Properties of ionic compounds
 - e. What is a crystal lattice and the arrangement.
 - f. What is metallic bonding?
4. CH 9: Covalent compounds
 - a. What elements make up covalent compounds?
 - b. Compare ionic vs covalent bonds. An e-neg difference of 1.7 is dividing factor.
 - c. What are pi bonds and sigma bonds?
 - d. Bond length Vs. Bond strength, and Bond dissociation energy.
 - e. What are molecules?
 - f. What are the common diatomic molecules?
 - g. Polar molecules, partial charges, what are they?
5. CH 10: Identify the types of chemical reactions
 - a. Be able to write the products for a double replacement and single replacement reactions.
 - b. Balancing equations
 - c. Writing equations from word names.
6. CH 11: The mole
 - a. Determine the molar mass of a compound
 - b. Calculate the amt of particles, ions, atoms, formula units, molecules in a compound.
 - c. Converting mole \rightarrow mole, mole \rightarrow grams, mole \rightarrow particles and visa versa
 - d. Converting grams \rightarrow mole, grams \rightarrow grams, grams \rightarrow particles visa vera
7. CH 12:
 - a. Stoichiometry conversions using balanced chemical equation
 - b. Mole \rightarrow mole, mole \rightarrow grams, grams \rightarrow grams

Objective Final Review For Chemistry

HESI



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Multi-objective Optimization: Techniques And Applications In Chemical Engineering (With Cd-rom) Gade Pandu Rangaiah,2008-12-22 Optimization has been playing a key role in the design planning and operation of chemical and related processes for nearly half a century Although process optimization for multiple objectives was studied by several researchers back in the 1970s and 1980s it has attracted active research in the last 10 years spurred by the new and effective techniques for multi objective optimization In order to capture this renewed interest this monograph presents the recent and ongoing research in multi optimization techniques and their applications in chemical engineering Following a brief introduction and general review on the development of multi objective optimization applications in chemical engineering since 2000 the book gives a description of selected multi objective techniques and then goes on to discuss chemical engineering applications These applications are from diverse areas within chemical engineering and are presented in detail All chapters will be of interest to researchers in multi objective optimization and or chemical engineering they can be read individually and used in one s learning and research Several exercises are included at the end of many chapters for use by both practicing engineers and students

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