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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$ now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively $\text{C}_6\text{H}_8\text{O}_6$ 2H^+ 2NO_2 $\text{C}_6\text{H}_6\text{O}_6$ $2\text{H}_2\text{O}$ 2NO choose 1 answer $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 a $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 b $\text{C}_6\text{H}_8\text{O}_6$ and NO c $\text{C}_6\text{H}_8\text{O}_6$ and H^+ d $\text{C}_6\text{H}_8\text{O}_6$ and H_2O

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