

2nd Year Honours BSc in Applied Science, University of Dundee
 CH101

This assignment will require you to use your knowledge of the following topics to answer the questions. You should refer to your lecture notes and any relevant textbooks. You should also refer to the relevant chapters in the following books: Principles of Chemistry by G. L. Fisher and Principles of Chemistry by R. A. Scott.

Questions

- 1) What is meant by **enthalpy of formation**? **10%**

Enthalpy of formation is the enthalpy change when one mole of a compound is formed from its elements in their standard states.
- 2) What is the difference between standard enthalpy of formation and standard enthalpy of combustion? **10%**

Standard enthalpy of formation is the enthalpy change when one mole of a compound is formed from its elements in their standard states. Standard enthalpy of combustion is the enthalpy change when one mole of a compound is completely burned in oxygen under standard conditions.
- 3) Explain, using appropriate equations, why the standard enthalpy of formation of carbon dioxide is more negative than that of carbon monoxide. **10%**

Standard enthalpy of formation of carbon dioxide is more negative than that of carbon monoxide because the combustion of carbon to carbon dioxide releases more energy than the combustion of carbon to carbon monoxide.
- 4) What are the standard enthalpy of formation and standard enthalpy of combustion of carbon monoxide? **10%**

Standard enthalpy of formation of carbon monoxide is $-110.5 \text{ kJ mol}^{-1}$. Standard enthalpy of combustion of carbon monoxide is $-283.5 \text{ kJ mol}^{-1}$.
- 5) What are the standard enthalpy of formation and standard enthalpy of combustion of carbon dioxide? **10%**

Standard enthalpy of formation of carbon dioxide is $-393.5 \text{ kJ mol}^{-1}$. Standard enthalpy of combustion of carbon dioxide is $-393.5 \text{ kJ mol}^{-1}$.
- 6) What does the sign of the standard enthalpy of formation of carbon monoxide tell you about the stability of carbon monoxide? **10%**

The negative sign of the standard enthalpy of formation of carbon monoxide indicates that carbon monoxide is more stable than its constituent elements.
- 7) What does the sign of the standard enthalpy of combustion of carbon monoxide tell you about the stability of carbon monoxide? **10%**

The negative sign of the standard enthalpy of combustion of carbon monoxide indicates that carbon monoxide is more stable than its constituent elements.
- 8) Define enthalpy of activation. State its units. Illustrate, by a graph, how enthalpy **10%**



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Jianjun Gao



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