

**Note-taking  
Worksheet****Atomic Structure and Chemical Bonds****Section 1 Why do atoms combine?**

- A. The \_\_\_\_\_, containing protons and neutrons, is at the center of an atom and is surrounded by the \_\_\_\_\_, an area of space around the \_\_\_\_\_ where \_\_\_\_\_ travel.
- \_\_\_\_\_ have a negative charge and do not travel in well-defined orbits.
  - Each element has a different \_\_\_\_\_, consisting of a particular number of protons, neutrons, and electrons.
- B. The \_\_\_\_\_ and \_\_\_\_\_ of electrons in the electron cloud determines the physical and chemical properties of the element.
- Electrons are arranged in different \_\_\_\_\_ at different distances from the nucleus.
  - The \_\_\_\_\_ an energy level is from the nucleus, the more electrons it can hold.
  - Electrons in the level closest to the nucleus have the \_\_\_\_\_ amount of energy; electrons farthest from the nucleus have the \_\_\_\_\_ amount of energy.
- C. Data from the \_\_\_\_\_ can be used to understand energy levels.
- The \_\_\_\_\_ is the same as the number of protons (or electrons) in an electrically neutral atom.
  - The number of electrons in an element's outermost energy level \_\_\_\_\_ from left to right across a period.
    - The first energy level is complete with two electrons, and the \_\_\_\_\_ has two elements.
    - The last element in each of the other \_\_\_\_\_ has \_\_\_\_\_ electrons in its outer energy level.
  - Each column in the periodic table contains one element family or group with \_\_\_\_\_ chemical properties.
    - The \_\_\_\_\_ in Group 8 do not combine easily with other elements because their energy levels are stable.

# Section 1 Reinforcement Structure Of The Atom

**National Institute of Standards and  
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