Section 18.2 Reversible Reactions and Equilibrium

OBJECTIVES:

- Describe how the amounts of reactants and products change in a chemical system at equilibrium.
- Identify three stresses that can change the equilibrium position of a chemical system.
- Explain what the value of Keq indicates about the position of equilibrium.

Reversible Reactions

- Some reactions do not go to completion as we have assumed
 - They may be ______ a reaction in which the conversion of reactants to products and the conversion of products to reactants occur simultaneously
- $2SO_2(g) + O_2(g) \rightarrow 2SO_2(g)$ = $2SO_2(g) + O_2(g) \leftarrow 2SO_2(g)$
- The two equations can be combined into one, by using a , which tells us that it is a reversible reaction:
- $2SO_2(g) + O_2(g) \leftrightarrow 2SO_3(g)$
- A chemical equilibrium occurs, and no net change occurs in the actual amounts of the components of the system.
- Even though the rates of the forward and reverse are equal, the
 of components on both sides may
 not be equal
 - An equlibrium position may be shown:
 B or A B
- 1% 99% 99% 1%
- Note the emphasis of the arrows direction
- It depends on which side is favored; almost all reactions are reversible to some extent

Le Chatelier's Principle

- The French chemist Henri Le Chatelier (1850-1936) studied how the equilibrium position <u>shifts</u> as a result of changing conditions
- Le Chatelier's principle:

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